

Thus, the oxidation number of sodium in the compound is $1+$ because one electron was transferred from the sodium atom. Because an electron is transferred to the chlorine atom, its oxidation number is $1-$.


- 7- In the chemical formula for any ionic compound, the symbol of the cation is always written first, followed by the symbol of the anion.
- 8- Subscripts, which are small numbers to the lower right of a symbol, represent the number of ions of each element in an ionic compound. If no subscript is written, it is assumed to be one.

Practice Problems **Chlorine** **Write Formulas**

Write formulas for the ionic compounds formed by the following ions.

19. potassium and iodide	21. aluminum and bromide
20. magnesium and chloride	22. sodium and nitride

23. **Challenge:** Write the general formula for the ionic compound formed by elements from the two groups shown on the periodic table at the right.



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Formulas for polyatomic ionic compounds:

- 9- polyatomic ions are ions made up of more than one atom. A polyatomic ion acts as an individual ion. Thus, the formula for a polyatomic compound follows the same rules used for a binary compound. Because a polyatomic ion exists as a unit, never change subscripts of the atoms within the ion.

Ionic Compounds Study Guide

Victor M. Corman



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