

The Hydrogen Atom – Student Guide

Background Material

Carefully read the background pages created through Eureka, Slides, and Experiments and answer the following questions in detail and understanding.

Question 1: Complete the following table which compares how the Bohr model and the Quantum model compare the Hydrogen atom. In each case, they both describe things in a unique way and in some ways they do not.

Concept Name	Bohr Model	Quantum Model
Electrons	positively charged particles in the atom	negatively charged particles in the atom
Orbitals	negatively charged and orbit the nucleus	negatively charged clouds
Location of the electron in the atom without	is in two fixed paths	electron cloud pattern
Energy of the electron in the atom without	electron has fixed energy levels	variable electron energy
Energy of the electron in the atom without	$E = -13.6 \text{ eV}$	$E = -13.6 \text{ eV}$

Question 2: Complete the following table which shows the comparison of two different theories by putting predictions made by each theory side-by-side. The table can be completed by using the Bohr model as a reference. Eureka, Slides, and Experiments are helpful resources, as well as your own understanding, intuition, hypothesis, and/or energy and the same methods through-out the entire course.

Hydrogen Atom Student Guide Naap Answers

Xiaolong Qi



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The Hydrogen Atom Energy Levels, 1976 *Student's Guide to Chemistry; a Modern Introduction* David Brooks, 1974

Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry Denise

Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts After a brief review of each section's critical ideas students are taken through stepped out worked examples try it yourself examples and chapter quizzes all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text *On the Energy Levels of a Model of the Compressed Hydrogen Atom*

S. R. de Groot, C. A. ten Seldam, 1946 Student Solutions Manual for General Chemistry John McMurry, Robert

Fay, 2013-03-07 By Joseph Topich Virginia Commonwealth University This manual for students contains solutions to selected all in chapter problems and even numbered end of chapter problems Student Study Guide and Solutions Manual to

accompany General Organic and Biological Chemistry, 1e Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond's text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry **Study Guide and Solutions Manual to Accompany Organic Chemistry, 11th**

Edition T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2013-03-25 This is the study guide and solutions manual to accompany Organic Chemistry 11th Edition **Student Study Guide/solutions Manual to Accompany Foundations of**

Inorganic, Organic, and Biological Chemistry Katherine J. Denniston, 1995 This student study guide has been written to accompany Foundations of Inorganic Organic and Biological Chemistry by Caret Denniston and Topping It was designed to complement the text not to be used in place of the text Each chapter of the study guide contains the following sections reorganized and expanded set of learning goals concise chapter summary in chapter solved problems list of key terms review problems Preface **The Hydrogen Atom** Rudolph M. Langer, California Institute of Technology. Division of Physics,

Mathematics and Astronomy, 1927 *The Reaction of Hydrogen Atom with Hydrogen Molecular Ion* Dion Allan Stams, 1985

The Two-dimensional Hydrogen Atom Todd S. Garon, 2011 *Chemistry* Steven S. Zumdahl, Thomas

Hummel, 2015-06 This manual provides detailed solutions for half of the end of chapter exercises designated by blue question numbers using the strategies emphasized in the text This manual has been thoroughly checked for precision and accuracy Answers to the For Review questions appear on the student website *Concerning the Theory of the Hydrogen Atom* Kaiser Schoen Kunz, 1937 *One-dimensional Hydrogen Atom* Kwan Lam Ng, 1990 The photoionization spectrum of highly excited hydrogen atoms in the energy range between E_0 and E_{ion} is examined in the presence of a

strong dc electric field Atomic hydrogen is excited from the ground state via a three photon process to high lying excited states Detailed experimental studies are made of the various properties of the resonant structures and the results are compared to numerical calculations It is found that the external electric field can be used to manipulate control and design specific atomic structures These resonant structures are nearly one dimensional and whose electronic distributions are highly extended along the field and which may have enormous electric dipole moments Such one dimensional hydrogen atoms are ideal tools for studying a number of physical phenomena and applications *Characterizations of Hydrogen Atom Systems* Aeronutronics Systems, Inc,1958

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