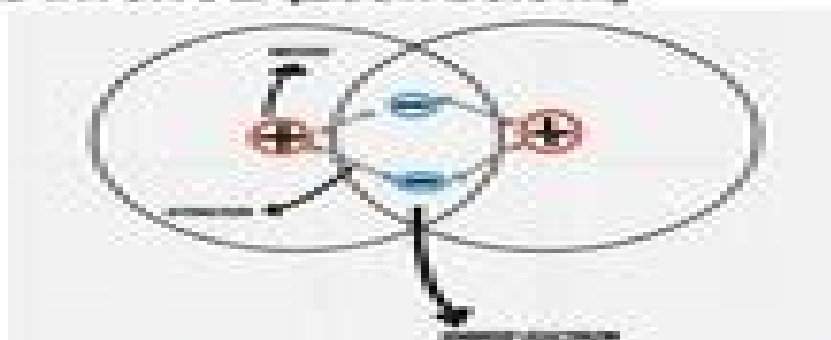


## Covalent bonding: (1/7)

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(This is when elements share electrons between each other.)

- 1) Covalent bonding makes full outer shells for every element.
- 2) Each covalent bond provides one extra shared electron for each other.
- 3) In covalent bonding there is also strong electrostatic attraction between the negatively charged shared electrons (the bonding pair) and the positively charged nuclei of the atoms involved. (Look below)



- 4) A plus sign (+) will attract a minus (-) for every element involved.
- 5) Regarding the bonds, count number of lines between each letter (like " $\text{N} = \text{O}$ ") this will say what type of bond it is. So " $\text{N} = \text{O}$ " is a double covalent bond.
- 6) A two covalent bond will look like " $\text{O} = \text{C} = \text{O}$ ".
- 7) When drawing covalent bonds, we use "dot and cross diagrams" to do this.

Remember: Count the number of electrons to double-check you have the right number in the outer shells.

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# Guided The Nature Of Covalent Bonding Answers

**Katrin Zwirgmaier**



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