

PLS1502

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May/June 2019

Introduction to African Philosophy

Duration : 2 Hours

100 Marks

EXAMINERS :

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SECOND :

MR N DLADLA

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product expression describes in mathematical terms the equilibrium

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web the chemical equation used was $Ca(OH)_2(s) \rightleftharpoons Ca^{2+}(aq) + 2OH^-(aq)$ the solubility product was determined to be $K_{sp} Ca^{2+} OH^-$ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

18.1 solubility product constant K_{sp} chemistry libretexts - Feb 01 2023

web jul 12 2023 because the concentration of a pure solid such as $Ca_3(PO_4)_2$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $K_{sp} = [Ca^{2+}]^3 [PO_4^{3-}]^2$

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get

the closest K_{sp} value to the theoretical K_{sp} for Ca was the determination through pH method 2 introduction calcium hydroxide is a strong base soluble in water

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