

GROUP 2 ELEMENTS AND THEIR COMPOUNDS

- show their outermost electron placed on s subshell
- readily react to lose 2 electrons
 - $\text{M} \rightarrow \text{M}^{2+} + 2\text{e}^-$

Trends down the group:

- 1^{st} ionisation energy decreases
 - Distance between outermost electron and nucleus increases
 - More shielding
 - Counteractive effect of increasing nuclear charge
- Atomic radius increases
- Melting point generally decreases
 - Much lower than all transition metals

Physical properties:

- Good conductors of heat and electricity
 - Sea of delocalised electrons
- Form white compounds
 - Transition metals form coloured compounds
- Oxidation number of +2
- All compounds are toxic or almost toxic
 - Be compounds show some distinct character because small highly charged cation polarises anion shell off anion
- Oxides / hydroxides are basic
 - React with acids (proton acceptors)
 - Example:
 - BeO
 - $\text{BeO} + 2\text{H}^+ \rightarrow \text{Be}^{2+} + \text{H}_2\text{O}$ which is amphoteric
- Density = 2.83 g/cm^3
 - Average density of transition metals is about 7 g/cm^3
- Compared to group 2 metals:
 - Harder + denser
 - Higher melting point
 - Metallic bond is stronger (higher charge on cation & more electrons delocalised)

Chemical properties:

- Strong reducing agents
- Oxidation is losing electrons
- Reduction is gaining electrons:
 - Reducing agent - donates electrons (reducing other substances)
 - Oxidising agent - accepts electrons (oxidising itself)
- E.g.
 - $\text{ZnO} + \text{CO} \rightarrow \text{Zn} + \text{CO}_2$
 - $\text{ZnO} \rightarrow \text{Zn}^{2+} + 2\text{e}^-$
 - $\text{CO} + 2\text{e}^- + 2\text{H}^+ \rightarrow \text{CO}_2$

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Table of Contents F321 Ocr A Chemistry May 2014 Paper

1. Understanding the eBook F321 Ocr A Chemistry May 2014 Paper
 - The Rise of Digital Reading F321 Ocr A Chemistry May 2014 Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying F321 Ocr A Chemistry May 2014 Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an F321 Ocr A Chemistry May 2014 Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from F321 Ocr A Chemistry May 2014 Paper

- Personalized Recommendations
 - F321 Ocr A Chemistry May 2014 Paper User Reviews and Ratings
 - F321 Ocr A Chemistry May 2014 Paper and Bestseller Lists
5. Accessing F321 Ocr A Chemistry May 2014 Paper Free and Paid eBooks
- F321 Ocr A Chemistry May 2014 Paper Public Domain eBooks
 - F321 Ocr A Chemistry May 2014 Paper eBook Subscription Services
 - F321 Ocr A Chemistry May 2014 Paper Budget-Friendly Options
6. Navigating F321 Ocr A Chemistry May 2014 Paper eBook Formats
- ePub, PDF, MOBI, and More
 - F321 Ocr A Chemistry May 2014 Paper Compatibility with Devices
 - F321 Ocr A Chemistry May 2014 Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of F321 Ocr A Chemistry May 2014 Paper
 - Highlighting and Note-Taking F321 Ocr A Chemistry May 2014 Paper
 - Interactive Elements F321 Ocr A Chemistry May 2014 Paper
8. Staying Engaged with F321 Ocr A Chemistry May 2014 Paper
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers F321 Ocr A Chemistry May 2014 Paper
9. Balancing eBooks and Physical Books F321 Ocr A Chemistry May 2014 Paper
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection F321 Ocr A Chemistry May 2014 Paper
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine F321 Ocr A Chemistry May 2014 Paper
- Setting Reading Goals F321 Ocr A Chemistry May 2014 Paper
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of F321 Ocr A Chemistry May 2014 Paper

- Fact-Checking eBook Content of F321 Ocr A Chemistry May 2014 Paper
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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