

## 82. Additional Questions

1. The percent by mass composition of a salt was found to be 76.58% potassium, 6.48% sulfur, and 16.94% oxygen. What is the empirical formula of this salt?

$$56.58g \text{ of K} \left( \frac{\text{molar K}}{39.098 \text{ g K}} \right) = 1.447 \text{ molar K} \quad \frac{1.447 \text{ molar K}}{0.723 \text{ molar S}} = 2$$

$$6.48g \text{ of S} \left( \frac{\text{molar S}}{32.065 \text{ g S}} \right) = 0.202 \text{ molar S} \quad \frac{0.202 \text{ molar S}}{0.202 \text{ molar S}} = 1$$

$$16.94g \text{ of O} \left( \frac{\text{molar O}}{16.000 \text{ g O}} \right) = 1.059 \text{ molar O} \quad \frac{1.059 \text{ molar O}}{0.202 \text{ molar S}} = 5$$

$$\text{empirical formula} = K_2S_1O_5$$

2. A compound consisting iron and sulfur was formed by combining 3.235 g of iron with 1.426 g of sulfur. What is the empirical formula of the compound?

$$3.235g \text{ of Fe} \left( \frac{\text{molar Fe}}{55.845 \text{ g Fe}} \right) = 0.05794 \text{ molar Fe} \quad \frac{0.05794 \text{ molar Fe}}{0.05794 \text{ molar Fe}} = 1 + 2 = 2$$

$$1.426g \text{ of S} \left( \frac{\text{molar S}}{32.067 \text{ g S}} \right) = 0.04447 \text{ molar S} \quad \frac{0.04447 \text{ molar S}}{0.05794 \text{ molar Fe}} = 0.7675 + 2 = 3$$

$$\text{empirical formula} = Fe_2S_3$$

3. Propylene has a molar mass of 42.08 g/mole and is composed of 84.3% hydrogen and 85.7% carbon. What is the molecular formula of propylene?

$$14.3g \text{ of H} \left( \frac{\text{molar H}}{1.008 \text{ g H}} \right) = 14.15 \text{ molar H} \quad \frac{14.15 \text{ molar H}}{7.134 \text{ molar C}} = 2$$

$$85.7g \text{ of C} \left( \frac{\text{molar C}}{12.011 \text{ g C}} \right) = 7.134 \text{ molar C} \quad \frac{7.134 \text{ molar C}}{7.134 \text{ molar C}} = 1$$

$$\text{empirical formula} = CH_2$$

$$CH_2 = 12.011 + 2(1.008) = 14.027 \quad \frac{42.08}{14.027} = 3$$

$$\text{molecular formula} = C_3H_6$$

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