



$$\Delta H_{298}^\circ = -68.73 \text{ kcal } (-287.57 \text{ kJ})$$

Atomic Weight of **B** = 10.82 grams.

Atomic Weight of **O** = 16.0 grams.

Required:

- (a) Standard heat of formation ΔH_{298}° of B_2O_3 in terms of per mole of B_2O_3 =?
- (b) Standard heat of formation ΔH_{298}° of B_2O_3 in terms of per gram of B_2O_3 =?

Solution:

Now, we will multiply above given equations with real numbers to make them equal and then we will add all equations with each other. It will yield the final equation which will lead us to the final solution of this problem.

1. Multiply eq (i) with 2.



Metallurgical Thermodynamics Problems Solution

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