

## ANSWERS TO TOPIC 1

### Review Questions

- |      |      |      |
|------|------|------|
| 1. 1 | 2. 3 | 3. 2 |
| 4. 4 | 5. 2 | 6. 2 |
| 7. 1 | 8. 3 | 9. 3 |
10. The plum pudding model pictures the atom with an even distribution of positive charges, with electrons throughout the atom. The planetary model pictures the nucleus at the center of the atom with electrons in orbits.
11. Rutherford fired positively charged alpha particles at gold foil. He expected them to pass through with little deflection. When some were greatly deflected, he reasoned that they must have been repelled by a massive positively charged particle, which was called the nucleus.
12. The Rutherford model pictured the electrons in fixed, orbiting positions. The wave-mechanical model pictures electrons as possessing both wave and particle properties.
- |       |       |       |
|-------|-------|-------|
| 13. 3 | 14. 3 | 15. 1 |
| 16. 3 | 17. 4 | 18. 4 |
| 19. 2 | 20. 4 | 21. 2 |
| 22. 1 | 23. 1 | 24. 3 |
| 25. 3 | 26. 2 | 27. 2 |
| 28. 3 | 29. 3 | 30. 2 |
| 31. 3 | 32. 4 | 33. 4 |
| 34. 4 | 35. 1 | 36. 2 |
| 37. 1 | 38. 1 | 39. 3 |
| 40. 1 | 41. 2 | 42. 3 |
| 43. 2 | 44. 4 | 45. 4 |
| 46. 3 | 47. 3 | 48. 4 |
| 49. 2 | 50. 3 | 51. 2 |
| 52. 2 | 53. 3 | 54. 1 |
| 55. 3 | 56. 3 |       |
57. Na  $1s^2 2s^2 2p^6 3s^1$   
Mg  $1s^2 2s^2 2p^6 3s^2$   
Al  $1s^2 2s^2 2p^6 3s^2 3p^1$   
Si  $1s^2 2s^2 2p^6 3s^2 3p^2$   
P  $1s^2 2s^2 2p^6 3s^2 3p^3$   
S  $1s^2 2s^2 2p^6 3s^2 3p^4$   
Cl  $1s^2 2s^2 2p^6 3s^2 3p^5$   
Ar  $1s^2 2s^2 2p^6 3s^2 3p^6$
- |       |       |       |
|-------|-------|-------|
| 58. 1 | 59. 4 | 60. 1 |
| 61. 1 | 62. 2 | 63. 4 |
| 64. 4 | 65. 2 | 66. 2 |
| 67. 2 | 68. 2 | 69. 1 |
70. The mixture contains at least three parts. One solid remained on the filter paper. The other two parts consist of a liquid with a solid dissolved in it.
71. If the calcium and sulfur can be separated by physical means, the substance is a mixture. If chemical means are required to separate them, it is a compound.
72. The picture should show several of each type of molecule evenly mixed.
73. A mixture of iron and oxygen would have the iron present as a solid and the oxygen present as a gas. The iron would be attracted to a magnet. A glowing splint would ignite into a flame if placed into the oxygen gas. If the iron and oxygen were

combined as a compound, it would be a solid that is not attracted to a magnet.

74. Pepperoni pizza is heterogeneous because individual parts can be seen.
75. A is a mixture of an element and a compound; one molecule contains identical atoms and the other one contains different types of atoms. B is a compound; all particles are identical and are made up of two different types of atoms. C is an element; all molecules are identical and each is made up of two identical atoms. D is a mixture of an element and two different compounds; one molecule contains identical atoms and the other two contain different types of atoms, but they differ from each other.
76. Container D is homogeneous because the particles are mixed on the molecular level.
77. The two substances are different materials because their composition varies.

### Regents Practice Questions

#### Part A

- |       |       |       |
|-------|-------|-------|
| 1. 1  | 2. 1  | 3. 1  |
| 4. 3  | 5. 1  | 6. 4  |
| 7. 2  | 8. 3  | 9. 3  |
| 10. 2 | 11. 3 | 12. 1 |
| 13. 2 | 14. 3 |       |

#### Part B

- |       |       |       |
|-------|-------|-------|
| 15. 4 | 16. 3 | 17. 1 |
| 18. 3 | 19. 1 | 20. 3 |
| 21. 1 | 22. 3 | 23. 2 |
| 24. 4 |       |       |

#### Part C

25. Because the two nuclei have different numbers of protons, they have different amounts of attraction for electrons. As a result, the amounts of energy separating the energy levels differ for the two elements.
26. The samples are not the same substance because the percent of each element differs from sample to sample.
27. Because he assumed the atom to have an even distribution of positive charges, he did not expect a concentration of positive charges that would strongly deflect or bounce back the alpha particles.
28. The spectra are based on the elements present in the compound. If some lines matched and some didn't, the compounds have a least one element in common and at least one differing element.
29. Column 1: 1, 2, 1; column 2: K, 39, 19; column 3: 52, 146, 92
30.  $(0.9778)(12) + (0.0222)(13) = 12.02$
31. They contain different numbers of neutrons.
32.  $(19.99 \times .909) + (20.99 \times 0.03) + (21.99 \times 0.88) = 20$
33.  $(12.00 \times 9.98) + (13.00 \times 1.07) = 12.00$
34. 100
35. Both isotopes contain 6 protons but different numbers of neutrons. Or: An isotope with a mass of 12.00 contains 6 protons and 6 neutrons, while the isotope with a mass of 13.00 contains 6 protons and 7 neutrons.

# June 2013 Global Regents Answer Key

**Hussin A.Rothana**



**June 2013 Global Regents Answer Key:**

This book delves into June 2013 Global Regents Answer Key. June 2013 Global Regents Answer Key is a vital topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into June 2013 Global Regents Answer Key, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:
  - Chapter 1: Introduction to June 2013 Global Regents Answer Key
  - Chapter 2: Essential Elements of June 2013 Global Regents Answer Key
  - Chapter 3: June 2013 Global Regents Answer Key in Everyday Life
  - Chapter 4: June 2013 Global Regents Answer Key in Specific Contexts
  - Chapter 5: Conclusion
2. In chapter 1, the author will provide an overview of June 2013 Global Regents Answer Key. This chapter will explore what June 2013 Global Regents Answer Key is, why June 2013 Global Regents Answer Key is vital, and how to effectively learn about June 2013 Global Regents Answer Key.
3. In chapter 2, the author will delve into the foundational concepts of June 2013 Global Regents Answer Key. This chapter will elucidate the essential principles that must be understood to grasp June 2013 Global Regents Answer Key in its entirety.
4. In chapter 3, the author will examine the practical applications of June 2013 Global Regents Answer Key in daily life. The third chapter will showcase real-world examples of how June 2013 Global Regents Answer Key can be effectively utilized in everyday scenarios.
5. In chapter 4, this book will scrutinize the relevance of June 2013 Global Regents Answer Key in specific contexts. The fourth chapter will explore how June 2013 Global Regents Answer Key is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, this book will draw a conclusion about June 2013 Global Regents Answer Key. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of June 2013 Global Regents Answer Key.

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