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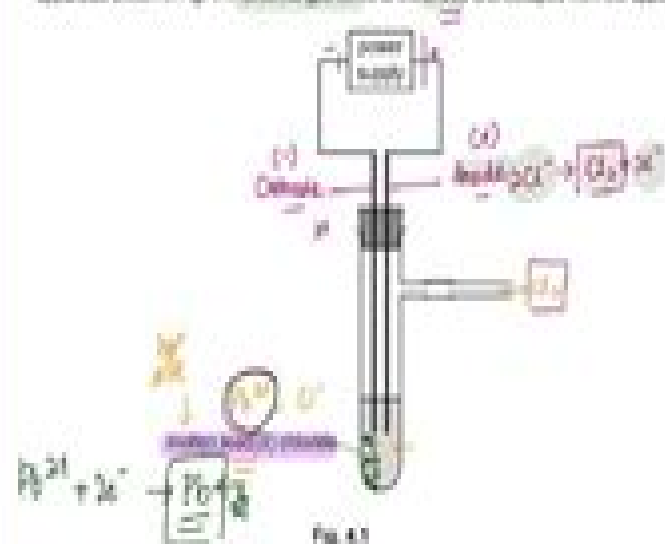
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(b) The student carries out an electrolysis experiment on molten lead(II) chloride using the apparatus shown in Fig. 4.1. (3) (Total gas taken at the anode and escapes from the apparatus.)



(i) Explain why lead(II) chloride needs to be molten before it will conduct electricity / react.
first moving electrons / mobile ions

(ii) Write the ionic half equation for the reaction occurring at the anode.



(iii) State the test for chlorine gas and describe the yellow fumes.

test Damp litmus paper

observations bleached / turns white

(iv) Describe what is observed at the cathode.

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Yicheng Fang

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Table of Contents Igcse Chemistry Paper Oct Nov

1. Understanding the eBook Igcse Chemistry Paper Oct Nov
 - The Rise of Digital Reading Igcse Chemistry Paper Oct Nov
 - Advantages of eBooks Over Traditional Books
2. Identifying Igcse Chemistry Paper Oct Nov
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Igcse Chemistry Paper Oct Nov
 - User-Friendly Interface
4. Exploring eBook Recommendations from Igcse Chemistry Paper Oct Nov
 - Personalized Recommendations

- Igcse Chemistry Paper Oct Nov User Reviews and Ratings
- Igcse Chemistry Paper Oct Nov and Bestseller Lists
- 5. Accessing Igcse Chemistry Paper Oct Nov Free and Paid eBooks
 - Igcse Chemistry Paper Oct Nov Public Domain eBooks
 - Igcse Chemistry Paper Oct Nov eBook Subscription Services
 - Igcse Chemistry Paper Oct Nov Budget-Friendly Options
- 6. Navigating Igcse Chemistry Paper Oct Nov eBook Formats
 - ePub, PDF, MOBI, and More
 - Igcse Chemistry Paper Oct Nov Compatibility with Devices
 - Igcse Chemistry Paper Oct Nov Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Igcse Chemistry Paper Oct Nov
 - Highlighting and Note-Taking Igcse Chemistry Paper Oct Nov
 - Interactive Elements Igcse Chemistry Paper Oct Nov
- 8. Staying Engaged with Igcse Chemistry Paper Oct Nov
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Igcse Chemistry Paper Oct Nov
- 9. Balancing eBooks and Physical Books Igcse Chemistry Paper Oct Nov
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Igcse Chemistry Paper Oct Nov
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Igcse Chemistry Paper Oct Nov
 - Setting Reading Goals Igcse Chemistry Paper Oct Nov
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Igcse Chemistry Paper Oct Nov
 - Fact-Checking eBook Content of Igcse Chemistry Paper Oct Nov

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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