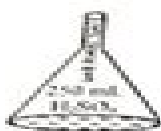


# অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার

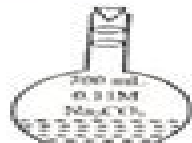


পটাসিয়াম কপ্লো, এখানে মোল পটাসিয়াম, আরও কপ্লোয়ন শীর্ষস্থানীয় কপ্লোয়নমুকের পটাসিয়াম এবং ময়েন টেম্পের পটাসিয়াম  
পটাসিয়াম উত্তর অধ্যায়ভিত্তিক মেয়েন হচ্ছে। এপুকের অনুশীলন করলে তুমি এ অধ্যায়টি থেকে মেয়েনমের পটাসিয়াম কপ্লোয়নমুকের  
এদের উত্তর লিখতে পারবে সহজেই।

প্রশ্ন ১



পাত্র-A



পাত্র-B

সি. কে. ১০। এর ক. ১/

- মোলের প্রবণ কী?
- "কেলস পটাসিয়াম লবণের প্রাকৃতিক, তাপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়"— ব্যাখ্যা করো।
- A-পাত্রে বিদ্যমান এসিডের ভর কত নির্ণয় করো।
- দ্রবণ দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অক্সিজেন দ্বারা অক্সিজেন হবে— বিশ্লেষণ করো।

১ নং প্রশ্নের উত্তর

1000 mL দ্রবণে কোন দ্রবের 1 mol দ্রবীভূত থাকলে উক্ত দ্রবকে মোলার দ্রবণ বলে।

কেলস পটাসিয়াম লবণের প্রাকৃতিক তাপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়। এর কারণ, তাপমাত্রা বৃদ্ধির সাথে সাথে প্রাকৃতিক বৃদ্ধি পেতে থাকলেও একটি নির্দিষ্ট তাপমাত্রার কেলস পটাসিয়াম লবণ নিষ্কৃতি হয়। তখন প্রাকৃতিক হ্রাস পেতে থাকে। উদাহরণস্বরূপ  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  এর কথা বলা যায়।  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  এর ক্ষেত্রে পটাসিয়াম দ্রবীভূত করার সময় বিশ্লেষণ প্রক্রিয়াটি তাপমাত্রা বৃদ্ধি 32.4°C তাপমাত্রার উপরে  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  যখন  $\text{Na}_2\text{SO}_4$  এ পরিণত হয় তখন বিশ্লেষণ প্রক্রিয়াটি তাপোৎপাদী হয়। এ কারণে 32.4°C তাপমাত্রার উপরে তাপমাত্রা বৃদ্ধিজনিত কারণে প্রাকৃতিক হ্রাস ঘটে।

১

আমরা জানি,  
 $w = MV$   
 $= (0.5 \times 98 \times 0.25) \text{ g}$   
 $= 12.25 \text{ g}$

অতএব, দ্রবীভূত  $\text{H}_2\text{SO}_4$  এর ভর  
 $= 12.25 \text{ g}$

উদাহরণের A-পাত্রে অবস্থিত,  
 $\text{H}_2\text{SO}_4$  এর আণবিক ভর,  
 $M = 1 \times 2 + 32 + 16 \times 4$   
 $= 2 + 32 + 64$   
 $= 98 \text{ g/mol}$

দ্রবণের ঘনমাত্রা,  
 $S = 0.5 \text{ M}$   
দ্রবণের আয়তন,  
 $V = 250 \text{ mL}$   
 $= \frac{250}{1000} = 0.25 \text{ L}$

দ্রবীভূত  $\text{H}_2\text{SO}_4$  এর ভর,  $w = ?$

১০ পাত্র-A:  $\text{H}_2\text{SO}_4$  এর মোল সংখ্যা  $= \left(0.5 \times \frac{250}{1000}\right) \text{ mol}$   
 $= 0.125 \text{ mol}$

পাত্র-B:  $\text{Na}_2\text{CO}_3$  এর মোল সংখ্যা  $= \left(0.11 \times \frac{700}{1000}\right) \text{ mol}$   
 $= 0.077 \text{ mol}$

সম্প্রতিক বিক্রিয়া:  $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$   
সমীকরণ হতে পাই, 1 mol  $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$

1 mol  $\text{Na}_2\text{CO}_3$  বিক্রিয়া করে  $\text{H}_2\text{SO}_4$  এর 1 mol এর সাথে  
 $0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol}$

অবশিষ্ট  $\text{H}_2\text{SO}_4$  এর পরিমাণ  $= (0.125 - 0.077) \text{ mol}$   
 $= 0.048 \text{ mol}$

পাত্র-A এর দ্রবণের সাথে পাত্র-B এর দ্রবণের বিক্রিয়ার ফলে 0.048 mol  $\text{H}_2\text{SO}_4$  অবশিষ্ট থাকায় মিশ্রিত দ্রবণের প্রকৃতি হবে অম্লীয়।

প্রশ্ন ২ একদল ছাত্র ল্যাবরেটরিতে X-পদ্ধতিতে 10 mg এর তরল পরিমাণ এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2%  $\frac{W}{V}$  HCl এর নমুনা তৈরি করে।

সি. কে. ১০। এর ক. ১/

- নোড কাকে বলে?
- $\text{Cl}_2$  অপোলার— ব্যাখ্যা করো।
- উদীপকের এসিডটির মোলার ঘনমাত্রা নির্ণয় করো।
- উদীপকের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো।

২ নং প্রশ্নের উত্তর

নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেকট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা সমান হবে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হবে সমযোজী বন্ধন বিশুদ্ধ ও অপোলার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোলার হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও C অপুতে তড়িৎ ঋণাত্মকতার পার্থক্য  $(3 - 3) = 0$ , তাই  $\text{Cl}_2$  বিশুদ্ধ সমযোজী ও অপোলার যৌগ। উল্লেখ্য যৌগিক অণুর ক্ষেত্রে এহুণ ঘাটক।

উদীপকে বর্ণিত এসিডটি হলো 2%  $\left(\frac{W}{V}\right) \text{ HCl}$

2%  $\left(\frac{W}{V}\right) \text{ HCl}$  এর অর্থ হলো—

100 mL দ্রবণে HCl এর ভর 2g

∴ 1000 mL " " " "  $\left(\frac{2 \times 1000}{100}\right) \text{ g}$   
 $= 20 \text{ g}$

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## **Table of Contents Hsc Chemistry 1st Paper Question 2solution**

1. Understanding the eBook Hsc Chemistry 1st Paper Question 2solution
  - The Rise of Digital Reading Hsc Chemistry 1st Paper Question 2solution
  - Advantages of eBooks Over Traditional Books
2. Identifying Hsc Chemistry 1st Paper Question 2solution
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Hsc Chemistry 1st Paper Question 2solution
  - User-Friendly Interface
4. Exploring eBook Recommendations from Hsc Chemistry 1st Paper Question 2solution
  - Personalized Recommendations
  - Hsc Chemistry 1st Paper Question 2solution User Reviews and Ratings
  - Hsc Chemistry 1st Paper Question 2solution and Bestseller Lists
5. Accessing Hsc Chemistry 1st Paper Question 2solution Free and Paid eBooks
  - Hsc Chemistry 1st Paper Question 2solution Public Domain eBooks
  - Hsc Chemistry 1st Paper Question 2solution eBook Subscription Services
  - Hsc Chemistry 1st Paper Question 2solution Budget-Friendly Options

6. Navigating Hsc Chemistry 1st Paper Question 2solution eBook Formats
  - ePub, PDF, MOBI, and More
  - Hsc Chemistry 1st Paper Question 2solution Compatibility with Devices
  - Hsc Chemistry 1st Paper Question 2solution Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Hsc Chemistry 1st Paper Question 2solution
  - Highlighting and Note-Taking Hsc Chemistry 1st Paper Question 2solution
  - Interactive Elements Hsc Chemistry 1st Paper Question 2solution
8. Staying Engaged with Hsc Chemistry 1st Paper Question 2solution
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Hsc Chemistry 1st Paper Question 2solution
9. Balancing eBooks and Physical Books Hsc Chemistry 1st Paper Question 2solution
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Hsc Chemistry 1st Paper Question 2solution
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Hsc Chemistry 1st Paper Question 2solution
  - Setting Reading Goals Hsc Chemistry 1st Paper Question 2solution
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Hsc Chemistry 1st Paper Question 2solution
  - Fact-Checking eBook Content of Hsc Chemistry 1st Paper Question 2solution
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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