বিষয় কোত :	1	7	6
Ø6:			7

হচনিৰ্বাচনি অভীকা

TO PROPERTY

ক্রবাদ পর

(২০২৩ সালের সিলেবাস অনুযায়ী)

সময়—২৫ মিনিট

বিলেৰ দুইবা । সৰবৰাহকৃত বছনিৰ্বাচনি অভীকাৰ উত্তৰপত্তে প্ৰস্তেই ক্ৰমিক নম্বৰে বিপত্তীতে প্ৰকল্প বৰ্ণসম্পদত বৃত্তসমূহ হতে স্তিক/সৰ্বোধকট উত্তৰে বৃত্ততি বল পছেন্ট কলম ছাৱা সম্পূৰ্ণ তথাট কৰ । প্ৰতিটি প্ৰস্তেৱ মান ১ ।]

#### অপ্লশতত কোনো প্ৰকাৰ দাশ/চিফ দেৱা যাবে দা।

- চিকিৎসা বিজ্ঞানে কিঞ্জিতখেরাপিতে কোনটি ব্যবহার করা হয়?
  - (P) X-ray

সিলেট বোর্ড

- B IR
- (9) MRI
- (M) UV

উদ্দীপক থেকে ২ এবং ৩ নং প্রপ্লের উত্তর দাও :

মৌল	বহিন্তে জনের ইলেবট্রন বিন্যাস			
A	(n-1)s <sup>1</sup>	Maked		
В	ns²np⁴	L decision		
С	(n+1)s <sup>2</sup> (n+1)p <sup>4</sup>	मान 2		

A, B এবং C প্রচলিত থৌলের প্রতীক বর।

- ২। 🗛 ৪ ৩ 🗛 C ভৌত অবস্থার ভিন্নতার ভারণ---
  - (ক) আছনিক বছন
  - (খ) সমযোগী বন্ধল
  - হাইছেলেন বছন
  - (খ) সন্নিবেশ বন্ধন
- ৩। A2B খৌগের ক্ষেত্র—
  - ে sp<sup>3</sup> সংকলণ ফটে
  - AL বছন কোপ 104-5°
  - iii. (Lp Lp) विकर्षन विनामान

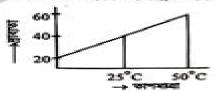
নিচের কোনটি সঠিক?

- (4) 14 H
- (4) II 4 III
- ৪। নিজের মৌলভলার ১ম আছনীকরণ বিভারের সঞ্জিক ক্রম কোনটিঃ
  - (\*) Be>B>N>O
    - N > O > Be > B
  - (\*f) O>N>B>Be (\*f) B>Be>N>O

- ৫। গোটনের প্রকৃত ভর কত।
  - (\*) 1-60×10<sup>-24</sup>g
  - (4) 1-66×10<sup>-24</sup>g
    - 9 1-673×10<sup>-24</sup>g
  - (v) 1-675×10<sup>-24</sup>g
- **৯। কোনটি প্রশন্ন অক্সাইড**া
  - (4) N<sub>2</sub>O<sub>3</sub>
- (4) NO<sup>3</sup>

12

- । S-ব্ৰক খৌলের সংখ্যা ককটিঃ
  - (<del>\*\*</del>) 7 (
- **9** 14 (%) 16
- শরমাপুর ৩য় শক্তিকরে ফুর্লনরত ইলেকট্রনের কৌশিক ভরবেশ এর মান কোনটিঃ
- 3h 2×
- (四) 经
- া কোনটি উভহবনী অপ্লাইত।
  - (1) KO<sub>2</sub>
- (4) Na<sub>2</sub>O<sub>2</sub> (4) Fe<sub>2</sub>O<sub>3</sub>
- উদীলক হতে ১০ এবং ১১ নং প্রপ্রের উত্তর দাও :



- ে। 200g NH<sub>4</sub>CI এর সম্পৃত প্রকাকে 50°C থেকে 25°C তালধারার ইত্তে করলে কর প্রাথ NH<sub>4</sub>CI কেলালিত হবে।
  - (m) 75

- (4) 50 (4) 90
  - /0 [भद्र भृक्षेत्र शुक्रेत

# **Hsc Chemistry 1st Paper Mcq**

V. K. Ahluwalia

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molar volume molecular ions molecular solids molecules moles Moseley law neutron properties non ideal behavior of gases orbital concept partial pressure calculations phase changes energies photons wave number Planck's quantum theory plasma state positive and negative ions pressure units properties of cathode rays covalent crystals properties of crystalline solids properties of positive rays quantum numbers quantum theory relative abundance Rutherford model of atom shapes of orbitals solid iodine structure solids properties solvent extraction spectrometer spin quantum number states of matter stoichiometry sublimation thermometry scales types of solids unit cell van der Waals equation vapor pressure and spectrum

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