

BOARD QUESTION PAPER: MARCH 2022

CHEMISTRY

Time: 3 Hrs

Max. Marks: 70

General Instructions:

The question paper is divided into four sections.

- (1) **Section A:** Q. No. 1 contains Ten multiple choice type of questions carrying One mark each.
Q. No. 2 contains Eight very short answer type of questions carrying One mark each.
- (2) **Section B:** Q. No. 3 to Q. No. 14 are Twelve short answer type of questions carrying Two marks each. (Attempt any Eight).
- (3) **Section C:** Q. No. 15 to Q. No. 26 are Twelve short answer type of questions carrying Three marks each. (Attempt any Eight).
- (4) **Section D:** Q. No. 27 to Q. No. 31 are Five long answer type of questions carrying Four marks each. (Attempt any Three).
- (5) Use of log table is allowed. Use of calculator is not allowed.
- (6) Figures to the right indicate full marks.
- (7) For each multiple choice type of question, it is mandatory to write the correct answer along with its alphabet, e.g. (a)...../(b)...../(c)...../(d)..... etc.

No marks(s) shall be given, if ONLY the correct answer or the alphabet of the correct answer is written. Only the first attempt will be considered for evaluation.

SECTION – A

Q.1. Select and write the correct answer for the following multiple choice type of questions:

[10]

- (i) The co-ordination number of atoms in body centred cubic structure (bcc) is _____.
(a) 4 (b) 6 (c) 8 (d) 12
- (ii) In calculating osmotic pressure, the concentration of solute is expressed in _____.
(a) molarity (b) molality
(c) mole fraction (d) percentage mass
- (iii) The enthalpy change for the chemical reaction $\text{H}_2\text{O}_{(s)} \longrightarrow \text{H}_2\text{O}_{(g)}$ is called enthalpy of _____.
(a) vapourisation (b) fusion
(c) combustion (d) sublimation
- (iv) Which of the following transition element shows maximum oxidation state?
(a) Sc (b) Fe (c) Mn (d) V
- (v) The correct formula for the complex compound, sodium hexacyanoferrate (III) is _____.
(a) $\text{Na} [\text{Fe}(\text{CN})_6]$ (b) $\text{Na}_2 [\text{Fe}(\text{CN})_6]$
(c) $\text{Na}_3 [\text{Fe}(\text{CN})_6]$ (d) $\text{Na}_4 [\text{Fe}(\text{CN})_6]$

Hsc Chemistry Question

Maureen Slade, Roger Slade



Hsc Chemistry Question:

HSC Chemistry Peter Lewis, Roger Michael Slade, 1986 **Excel HSC Chemistry** Jim Stamell, 2011 ISBN 9781741252996 AUTHOR Jim Stamell RRP 39 95 PAGES 428 pp SPECIFICATION Softcover perfect bound 280 mm x 210 mm STATUS New edition PUBLICATION DATE April 2008 The EXCEL HSC Chemistry guide is directly linked to the syllabus with every single dot point of the HSC Chemistry syllabus appearing in the margin of the book You can write in the guide so your study is focused and your notes are structured This guide comes in a brand new format that makes even better use of your study time up to date coverage of the core topics plus 3 Option topics Industrial Chemistry Shipwrecks Corrosion and Conservation and Forensic Chemistry this guide is organised just like the HSC syllabus so the students learn to section the theoretical part is under routine headings and the students section the practical part is under headings like First hand Second hand and Investigations and Problem Solving %this way you will be able to see at a glance what the theoretical and practical work is all main headings in each chapter 1 1 2 1 etc are directly from the syllabus word for word %this way you can easily match the Excel guide to the syllabus an alphabetical list of all the key definitions and concepts you should know from each chapter %an efficient way of learning all the definitions in one go chapter syllabus checklist with every single dot point listed in checklist form for each chapter %a fantastic way of testing that you know all the work hundreds of key concept questions with answers %questions that test you recall of knowledge in each chapter HSC type questions for every section in each chapter with clock icons to tell you how much time you will have to answer the questions in the HSC %this way you can test yourself on HSC type questions under HSC type time pressure an examiner maximiser feature ticks to show the mark distribution and answers to all HSC type questions %all you need to answer HSC type questions two sample HSC papers with an examiner maximiser feature plus answers %not one but two up to date sample papers the Excel syllabus summary notes a detachable section at the end of the guide where every single dot point of each chapter is summarised for you% a comprehensive and compact summary of the whole course in 32 pages **Cambridge Checkpoints HSC Chemistry 2017-19** Maureen Slade, Roger Slade, 2016-06-29 Excel HSC Chemistry C. M. Roebuck, 2003 **Past HSC Chemistry Questions and Worked Answers 1986-1997 by New Syllabus Topics** Aaron Butler, 1998 **Excel Revise in a Month HSC Chemistry** Will Marchment, 2007 **Excel HSC Maths Topic by Topic Exam Questions** J. Compton, Allyn Jones, 2000 Provides an essential supplement to the core Maths study guide with extra practice working through exam questions for complete exam preparation HSC Chemistry Marilyn Schell, Margaret Hogan, 2007-09-01 In this book you will find nearly 600 typical multiple choice examination questions nearly 600 extension questions plus summaries for the Year 12 Chemistry course Answers to all questions are provided Questions follow the dot points in the Board of Studies syllabus and focus on the three core topics plus the option topic Shipwrecks Corrosion and Conservation Excel Chemistry C. M. Roebuck, 1995 **Excel HSC Maths Extension 1** S. K. Patel, 2005 This comprehensive study guide covers the

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