

Question 33 (4 marks)

A student adds 1.17 g of $\text{Al}(\text{OH})_3(s)$ to 0.500 L of $0.100 \text{ mol L}^{-1} \text{HCl}(aq)$.

4

Calculate the pH of the resulting solution. Assume that the volume of the resulting solution is 0.500 L.

① Write a balanced chemical equation



② Find no. of moles of OH^- and H^+

$$n(\text{Al}(\text{OH})_3) = \frac{m}{M_r} = \frac{1.17}{26.98 + 3(16 + 1.008)} = 0.01499 \dots \text{mol}$$

$$n(\text{OH}^-) = 3 \times n(\text{Al}(\text{OH})_3) = 0.04499 \dots \text{mol}$$

$$n(\text{HCl}) = c \times V = 0.5 \times 0.1 = 0.05 \text{ mol}$$

$$n(\text{H}^+) = n(\text{HCl}) = 0.05 \text{ mol}$$

③ Determine which reagent is in excess and how much

As $n(\text{H}^+) > n(\text{OH}^-)$ $\therefore \text{H}^+$ is in excess

$$n(\text{H}^+)_{\text{excess}} = 0.05 - 0.04499 \dots = 0.00500 \dots \text{mol}$$

④ Calculate the $[\text{H}^+]$ and thus find pH.

$$[\text{H}^+] = \frac{n}{V} = \frac{0.00500 \dots}{0.500} = 0.0100 \dots \text{mol L}^{-1}$$

$$\text{pH} = -\log_{10} [\text{H}^+] = 2.000 \text{ (3 dp)}$$

Using d.p. 0015 f for pH

Please turn over

Hsc Chemistry 6 User Guide

Albert A Gayle



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of a symposium sponsored by the Hydrometallurgy and Electrometallurgy Committee and the Materials Characterization Committee of the Extraction and Processing Division of TMS The Minerals Metals Materials Society Held during the TMS 2012 Annual Meeting Exhibition Orlando Florida USA March 11 15 2012 **Excel HSC Biology** Diane Alford, Jennifer Hill, 2008

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