Basic Stoichiometry

Wilty ? In this activity we will address the question:

Hore do I convert between nucles of different chemical species in a given reaction?

Model I

2A + 3B -> 5C + 4D

2 mol A produces 5 mol C 3 mol B produces 4 mol D

2 mail A reacts with 3 mol B.

4 mol A produces 10 mol C 6 mol B produces 8 mol D

6 mol A means with 9 mel B

ii. What do the coefficients of a balanced equation represent?

 If 15 mol B are reacted, how many real C are formed? Show your work.

 If 6 mail A are reacted, how many and D are formed? Explain IN WORDS how you figured this our. If 8 and of A are reacted, how many real B are used? Show year work.

Model 2

 $2A + 3B \Rightarrow 3C + 4D$.

If 3 mol A react, how many mol C are produced?

3 mail A * 3 mail C = 7.5 mol C

In general: mol known t <u>mol anknown</u> = mol unknown mol known

 If 21.0 mol A are reacted, how many mail C are produced? Using dimensional analysis, alone how you calculated your asswer. If 21,0 mol B are reacted, how many mol A most have reacted?

Hs Chemistry Pogil Activity Basic Stoichiometry Answers

David E. Newton

Hs Chemistry Pogil Activity Basic Stoichiometry Answers:

Calculating in Chemistry Cosmas C Mbuzi, 2022-09-20 This book is intended to help students fully grasp calculations involving reacting quantities Students in higher courses may find it a helpful revision and enhance their clarity The book completely discusses the key topics in basic stoichiometry including the mole concept reacting quantities and empirical and molecular formulas It begins with the basic concepts and formulas required to covert various quantities to moles or amount of substance This particularly useful to a beginner Chapter 2 describes how to calculate reacting quantities and therefore provides a general step by step framework or approach by which to solve these problems The chapter also describes and applies the concept of a limiting reagent Two methods of determining a limiting reagent are explained and illustrated The concept of limiting reagent is extended to reacting volumes of gases with a short cut method A short cut method for solving reacting quantities involving masses and volumes of gases is also given Chapter 3 describes the calculations involved in the practical determination of molecular and empirical formulas A clear meaning of percentage composition of mass is provided and used to solve problems in a step by step manner In chapter 4 we discuss percentage yield and purity A number of examples are given to illustrate how formulas of yield and purity are used in various circumstances A student will find these examples helpful in relating different formulas for percentage purity The last chapter introduces a graphical method for reacting quantities The method may provide a new way of looking at chemical reactions Examples are given to illustrate the method including how it can be used to determine limiting reagents It is hoped that the book will provide all the necessary knowledge and skills to students studying an introductory chemistry course Teachers may also find this book a good resource for their lessons in stoichiometry
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representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in con trasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction

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