# 97113

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: B.E./B.Tech. DEGREE EXAMINATION, NOV/DEC 2014.

Benneh.

: Mechanical Engineering (Taxeil Medium)

Semester

: 111 - R2013

Code No./Subject

: ME 6301 - ENGINEERING THERMODYNAMICS

# KEY

 When a system satisfies all modes of equilibrium, i.e., chemical, thermal and mechanical equilibriums, then the system is said to be in Thermodynamic equilibrium.

2.

- 1. Both are recognized only as they cross the boundary of a system. Thus is, both are boundary phenomena.
- Systems possess energy, but not heat or work. That id, beat and work are transient phenomena.
- Both are associated with a process, not with a state. Unlike properties, heat and work have no meaning at a state.
- 4. Both are park functions. That is, their magnitudes depend on the path followed during a process as well as on the end states.

3.

Solution Since the engine is supposed to produce 60 [kW] of work while absorbing 100 [kW] of heat, the efficiency is 60(100 = 0.6). Now, a reversible engine working between the same reservoirs will have an efficiency of  $\eta = 1 - (T_1/T_1) = 1 - (500/1000) = 0.5$ . Then, it is clear that the proposed engine is more efficient than a reversible one. Since, by the Carnot theorem, this is impossible we note that this engine cannot be constructed and investment in its development will be a waste of resources.

4.

Solution A turbine is a flow system with the default assumptions of steady state operation  $(\hat{E}_{Cr} = 0 - \hat{S}_{Cr})$ , steady adiabatic flow  $(\hat{m}_{Cr} = 0 - \hat{Q})$ , and negligible changes in the kinetic and potential energies  $(\Delta e_i = 0 - \Delta e_j)$ .

1. Then, the relation for availability is

$$A = B_s = \dot{B}_s = \dot{m} \cdot [(h_s - T_0 s_s) - (h_s - T_0 s_s)] = \dot{m} \cdot [(h_t - h_s) - T_0 \cdot (s_t - s_s)]$$

- 2. Since the flow is isentropic,  $s_i = s_s$  and the relation for availability reduces to,  $\hat{\beta}_i = \frac{1}{2\pi} \cdot [(\delta_i h_s)]$ .
- From the steam tables, at 7 [bar] and 250 [°C], h<sub>i</sub> = 2954.0 [kJ/kg] and s<sub>i</sub> = 7.1066 [kJ/kg·K] = s<sub>i</sub> (since the process is isentropic). Then from the standard formulae, s<sub>i</sub> = 0.9582 and h<sub>i</sub> = 2581.0 [kJ/kg].
- Substituting the values in the above fermula for availability and evaluating it gives,
   A = (5) · (2954.0 2581.0) = 1865 [kW].

# **Hnd Mechanical Thermodynamic Exam Paper 2014**

Kevin Jost, Corina Sandu, Ryan Gehm

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