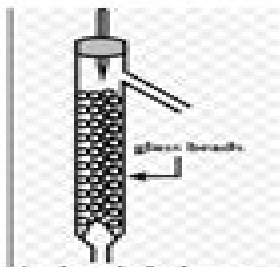


Suggest why chloroethanoic acid is stronger than ethanoic acid	(this question is already discussed a little further down.)		
fill in rate table	1.00 0.16 1.35 I beg someone write what the values were if they remember	K=250 Yes i agree - needs to be 1.4 as all values were to two sig figs Weren't we they to 3 dp though? The rate of the first one was 1.00 so I'm guessing 3sf of 2dp	3
empirical formula	C6H10O3S2	Percentages dudbt add up to 100 fuck me Didn't even see the sulphur, fml same same i thought i meant i was in the fuel area. YAY i got this!!	3
the labeling distillation	for the tube it was fractionating column or glass beads the bottom of the condenser was water in and top was water out	 <p>reject anti bumping granules might accept silica BEADS</p>	3
Draw a line for the reaction with chloroethanoic acid Suggest why chloroethanoic acid is a stronger acid than ethanoic acid.	<p>Was the line of chloroethanoic acid to the left of the first line, but levelled off at the same point? Yes but it levelled out before, same height though. Ah ok.</p> <p>Dissociates more Chlorine is more electronegative than hydrogen so will withdraw electron density from the carboxyl group meaning that the O-H bond breaks more easily dissociating into H⁺ making the solution more acidic</p>	Cl is electron withdrawing. it polarises the bond, weakens the O-H bond which releases more H ⁺ than ethanoic acid Neg inductive effect? Chloroethanoic acid forms a more stable conjugate base?	3
NMR Question?	<p>Use the integration values and the ratios to identify the structure at two peaks.</p> <p>Draw this structure.</p>	Was there a typo though? Regarding what? No dont think so, it was just because 2 of the peaks were the same environment almost so people got thrown off	5 2 1 2

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Terry C. Jones



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