F325 Past paper Answers

- (a) G: CO (1) HCOOH/H₂SO₄ → CO + H₂O (1)
 - (b) H: C (1) $C_{12}H_{22}O_{11} \rightarrow 12C + 11H_2O$ (1)
 - (c) I $C_4H_6O_2$ (1) $2C_2H_4O_2 \rightarrow C_4H_6O_2 + 2H_2O$ (1) Structure:



- (a) empirical formula
 N: O = 63.64/14: 36.36/16 (1)
 = 4.56: 2.27 = 2: 1. Empirical formula = N₂O (1)
 molecular formula
 Af, of gas = 1.833 × 24 = 44 (1)(calc 43.992)
 with these two pieces of evidence, assume that molecular formula = N₂O
 - (b) any chemical that reacts to produce gas:
 e.g. carbonate and CO₂ (1)
 accept: metal more reactive than Pb and H₂
 balanced equation to match chemical added (1)
 - (c) $M_{\bullet}(\text{Lidocaine}) = 236 \text{ (1)}$ moles Novocaine = $100 \times 10^{-3}/236 = 4.24 \times 10^{-4} \text{ (1)}$ concentration of Novocaine = $4.24 \times 10^{-4} \times (1000/5)$ = $0.0847/0.0848 \text{ mol dm}^{-3} \text{ (1)}$
 - (d) mass $C = 12 \times 3.74/44.0 = 12 \times 0.085 = 1.02 \text{ g (1)}$ mass $H = 2/18 \times 0.918 = 0.102 \text{ g (2)}$ mass O = 1.394 = (1.020 + 0.102) = 0.272 gratio C : H : O = 1.02/12 : 0.102/1 : 0.272/16 (1) = 0.0850 : 0.102 : 0.0170 / 5 : 6 : 1 / C₃H₆O (1) $C_3H_{10}O$ has relative mass of 82 M_6 is $164 = 2 \times 82$ M_7 molecular formula $= C_{10}H_{12}O_2 \text{ (1)}$
- 3. (i) $4PCl_5 + 10MgO \rightarrow P_4O_{10} + 10MgCl_2$ (1) 1 $\frac{100}{2}$
 - (ii) 100g P₄O₁₀ = 284 / 0.35(2) mol (1) moles PCl₅ needed = 4 × 0.352 = 1.408/1.4 mol (1) mass PCl₅ = 1.4(08) × 208.5 = 293.568 / 294 g/ 291.9 g (1)

2

3

3

5

[7]

[13]

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Kathleen Armour



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