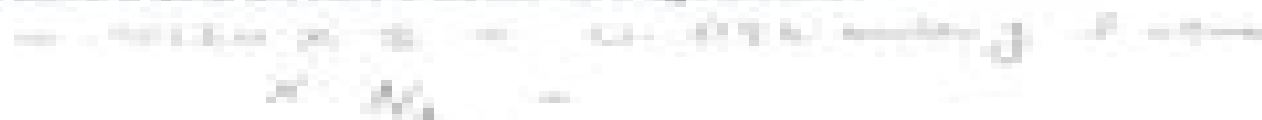


100) Carbonic acid reacts as a buffer at pH 7.4. How many moles of H_2CO_3 react with 1 mole of HCO_3^- to form 1 mole of H_2O and 1 mole of CO_2 ?



Answer: $1 \text{ mole H}_2\text{CO}_3$ reacts with 1 mole HCO_3^- to form $1 \text{ mole H}_2\text{O}$ and 1 mole CO_2 .

101) The only intracellular buffer in body fluids is bicarbonate.

102) Describe how the bicarbonate buffer system works.

The bicarbonate buffer system works by converting CO_2 and H_2O into H_2CO_3 and HCO_3^- . H_2CO_3 is a weak acid and HCO_3^- is a weak base. They form a buffer system that can resist changes in pH.

103) Describe how the bicarbonate buffer system works in the blood.

In the blood, CO_2 and H_2O are converted to H_2CO_3 and HCO_3^- . H_2CO_3 is a weak acid and HCO_3^- is a weak base. They form a buffer system that can resist changes in pH.

104) The bicarbonate buffer system is a component of the body's acid-base balance.

105) Describe how the bicarbonate buffer system works in the blood.

In the blood, CO_2 and H_2O are converted to H_2CO_3 and HCO_3^- .

Answer: $1 \text{ mole H}_2\text{CO}_3$ reacts with 1 mole HCO_3^- to form $1 \text{ mole H}_2\text{O}$ and 1 mole CO_2 .



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