

**Lewis Structure Worksheet #2****Key**

Draw Lewis Structures for each of the following molecules or ions:

|                                                                                                                                                                                                                                                                               |                                                                                                                                                                                                                                                                                                                                                                                |
|-------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|--------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|
| 1. $\text{Br}_2$<br>14<br>$:\ddot{\text{Br}} - \ddot{\text{Br}}:$                                                                                                                                                                                                             | 2. $\text{BH}_3$<br>6<br>$\begin{array}{c} \text{H} - \text{B} - \text{H} \\   \\ \text{H} \end{array}$                                                                                                                                                                                                                                                                        |
| 3. $\text{O}_2$<br>12<br>$:\ddot{\text{O}} = \ddot{\text{O}}:$                                                                                                                                                                                                                | 4. $\text{PS}_3$<br>24<br>$\begin{array}{c} :\ddot{\text{S}} = \text{P} - \ddot{\text{S}}: \\   \\ :\ddot{\text{S}}: \end{array} + \text{resonance}$                                                                                                                                                                                                                           |
| 5. $\text{H}_2\text{CS}$<br>12<br>$\begin{array}{c} \text{H} \\ \diagdown \\ \text{C} = \ddot{\text{S}}: \\ \diagup \\ \text{H} \end{array}$                                                                                                                                  | 6. $\text{CHCl}_3$<br>26<br>$\begin{array}{c} :\ddot{\text{Cl}}: \\   \\ \text{H} - \text{C} - \ddot{\text{Cl}}: \\   \\ :\ddot{\text{Cl}}: \end{array}$                                                                                                                                                                                                                       |
| 7. $\text{SF}_2$<br>20<br>$:\ddot{\text{F}} - \ddot{\text{S}} - \ddot{\text{F}}:$                                                                                                                                                                                             | 8. $\text{SHF}$<br>14<br>$\text{H} - \ddot{\text{S}} - \ddot{\text{F}}:$                                                                                                                                                                                                                                                                                                       |
| 9. $\text{N}_2\text{F}_4$<br>38<br>$\begin{array}{c} :\ddot{\text{F}}: \quad :\ddot{\text{F}}: \\   \quad   \\ :\ddot{\text{N}} - \ddot{\text{N}}: \\   \quad   \\ :\ddot{\text{F}}: \quad :\ddot{\text{F}}: \end{array}$                                                     | 10. $\text{XeO}_4$<br>32<br>$\begin{array}{c} :\ddot{\text{O}}: \quad :\ddot{\text{O}}: \\ \diagdown \quad \diagup \\ \text{Xe} \\ \diagup \quad \diagdown \\ :\ddot{\text{O}}: \quad :\ddot{\text{O}}: \end{array}$                                                                                                                                                           |
| 11. $\text{N}_2\text{O}_3$ (dinitrogen trioxide)*<br>28e<br>$\begin{array}{c} :\ddot{\text{O}} = \ddot{\text{N}} - \text{N} = \ddot{\text{O}}: \\ \diagdown \quad \diagup \\ :\ddot{\text{O}}: \end{array} + \text{resonance}$                                                | 12. $\text{NO}_2$ (nitrogen dioxide)*<br>17e<br>$:\ddot{\text{O}} = \ddot{\text{N}} - \ddot{\text{O}}: + \text{resonance}$                                                                                                                                                                                                                                                     |
| 13. $\text{N}_2\text{O}_4$ (dinitrogen tetraoxide)*<br>34<br>$\begin{array}{c} :\ddot{\text{O}}: \quad :\ddot{\text{O}}: \\ \diagdown \quad \diagup \\ \text{N} - \text{N} = \ddot{\text{O}}: \\ \diagup \quad \diagdown \\ :\ddot{\text{O}}: \end{array} + \text{resonance}$ | 14. $\text{N}_2\text{O}_5$ (dinitrogen pentaoxide)*<br>40e<br>$\begin{array}{c} :\ddot{\text{O}}: \quad :\ddot{\text{O}}: \quad :\ddot{\text{O}}: \\ \diagdown \quad \diagup \quad \diagdown \quad \diagup \\ \text{N} - \ddot{\text{O}} - \text{N} = \ddot{\text{O}}: \\ \diagup \quad \diagdown \\ :\ddot{\text{O}}: \quad :\ddot{\text{O}}: \end{array} + \text{resonance}$ |

\*Problems 11-14 are more challenging and may have more than one correct structure. Please note that if a molecule has an odd number of valence electrons, at least one atom in the molecule will NOT fulfill the octet rule.

# Lewis Structure Worksheet Guided Practice Answers

**J Dewey**



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