

1 This question is about reversible reactions.

Sulfur trioxide is produced in a reversible reaction.

The equation for the reaction is:



1.1 What is the correct balanced equation for the reaction?

Tick **one** box.

- ☐ $2\text{SO}_2 + \text{O}_2 \rightleftharpoons \text{SO}_3$
- ☐ $\text{SO}_2 + 2\text{O} \rightleftharpoons \text{SO}_3$
- ☐ $2\text{SO}_2 + \text{O}_2 \rightleftharpoons 2\text{SO}_3$
- ☐ $2\text{SO}_2 + 2\text{O}_2 \rightleftharpoons 2\text{SO}_3$

[1 mark]

1.2 The reaction of sulfur dioxide and oxygen to form sulfur trioxide is exothermic.

Which of the following statements is true?

Tick **one** box.

- ☐ The activation energy for the reaction of sulfur dioxide with oxygen is negative.
- ☐ Sulfur trioxide is colder than sulfur dioxide and oxygen.
- ☐ Heating the reaction will cause more sulfur trioxide to form.
- ☐ The reaction of sulfur trioxide to form sulfur dioxide and oxygen is endothermic.

[1 mark]

June 2013 Paper 4chemistry

Tao Wei

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name Tao Wei.

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Table of Contents June 2013 Paper 4chemistry

1. Understanding the eBook June 2013 Paper 4chemistry
 - The Rise of Digital Reading June 2013 Paper 4chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying June 2013 Paper 4chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an June 2013 Paper 4chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from June 2013 Paper 4chemistry
 - Personalized Recommendations
 - June 2013 Paper 4chemistry User Reviews and Ratings

- June 2013 Paper 4chemistry and Bestseller Lists
- 5. Accessing June 2013 Paper 4chemistry Free and Paid eBooks
 - June 2013 Paper 4chemistry Public Domain eBooks
 - June 2013 Paper 4chemistry eBook Subscription Services
 - June 2013 Paper 4chemistry Budget-Friendly Options
- 6. Navigating June 2013 Paper 4chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - June 2013 Paper 4chemistry Compatibility with Devices
 - June 2013 Paper 4chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of June 2013 Paper 4chemistry
 - Highlighting and Note-Taking June 2013 Paper 4chemistry
 - Interactive Elements June 2013 Paper 4chemistry
- 8. Staying Engaged with June 2013 Paper 4chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers June 2013 Paper 4chemistry
- 9. Balancing eBooks and Physical Books June 2013 Paper 4chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection June 2013 Paper 4chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine June 2013 Paper 4chemistry
 - Setting Reading Goals June 2013 Paper 4chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of June 2013 Paper 4chemistry
 - Fact-Checking eBook Content of June 2013 Paper 4chemistry
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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