

- 1 Use your copy of the periodic table to help you answer these questions.
- Predict the formula of each of the following compounds.
 - barium oxide
 - boron oxide
 - Give the formula of the following ions.
 - sulfide
 - gallium
 - Draw a diagram showing the arrangement of the valency electrons in one molecule of the covalent compound nitrogen trichloride. Use x to represent an electron from a nitrogen atom. Use o to represent an electron from a chlorine atom.
 - Potassium and vanadium are elements in Period IV.
 - State two differences in their physical properties.
 - Give two differences in their chemical properties.
 - Fluorine and astatine are halogens. Use your knowledge of the other halogens to predict the following:
 - the physical state of fluorine at r.t.p.
 - the physical state of astatine at r.t.p.
 - two similarities in their chemical properties.

Cambridge IGCSE Chemistry 0620 Paper 3 Q4 June 2007

- 2 Across the world, food safety agencies are investigating the presence of minute traces of the toxic hydrocarbon, benzene, in soft drinks.

It is formed by the reduction of sodium benzoate by vitamin C.



- Sodium benzoate is a salt. It has the formula C_6H_5COONa . It can be made by neutralising benzoic acid using sodium hydroxide.
 - Deduce the formula of benzoic acid.
 - Write a word equation for the reaction between benzoic acid and sodium hydroxide.
 - Name two other compounds that would react with benzoic acid to form sodium benzoate.

- Benzene contains 92.3% of carbon and its relative molecular mass is 78.
 - What is the percentage of hydrogen in benzene?
 - Calculate the ratio of moles of C atoms: moles of H atoms in benzene.
 - Calculate its empirical formula and then its molecular formula.
- This shows the structural formula of Vitamin C.
 - What is its molecular formula?
 - Name the two functional groups that are circled.



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- 3 The following is a list of the electron distributions of atoms of unknown elements.

element	electron distribution
A	2,6
B	2,8,4
C	2,8,8,2
D	2,8,18,8
E	2,8,18,8,1
F	2,8,18,18,7

- Choose an element from the list for each of the following descriptions.
 - It is a noble gas.
 - It is a soft metal with a low density.
 - It can form a covalent compound with element A.
 - It has a giant covalent structure similar to diamond.
 - It is a diatomic gas with molecules of the type X_2 .
- Elements C and A can form an ionic compound.
 - Draw a diagram that shows the formula of this compound, the charges on the ions, and the arrangement of the valency electrons around the negative ion. Use o to represent an electron from an atom of C. Use x to represent an electron from an atom of A.
 - Predict two properties of this compound.

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