

- (ix) The standard potential of the cell in the following reaction is _____
 $\text{Cd}_{(\text{s})} + \text{Cu}_{(\text{aq})}^{2+} \longrightarrow \text{Cd}_{(\text{aq})}^{2+} + \text{Cu}_{(\text{s})}$
 ($E_{\text{Cd}}^{\circ} = -0.403 \text{ V}$, $E_{\text{Cu}}^{\circ} = 0.334 \text{ V}$)
 (a) -0.737 V (b) 0.737 V
 (c) -0.069 V (d) 0.069 V
- (x) The value of $[\text{H}_3\text{O}^+]$ in mol lit^{-1} of 0.001 M acetic acid solution ($K_a = 1.8 \times 10^{-5}$) is _____
 (a) 1.34×10^{-1} (b) 1.34×10^{-2}
 (c) 1.34×10^{-3} (d) 1.34×10^{-4}

Q.2. Answer the following questions:

[8]

- Write the product formed when alkyl halide reacts with silver nitrite.
- Write the name of product formed, when acetone is treated with 2, 4-dinitrophenyl hydrazine.
- Write the name of biodegradable polyamide copolymer.
- Identify the molecularity of following elementary reaction:
 $\text{NO}_{(\text{g})} + \text{O}_{(\text{g})} \longrightarrow \text{NO}_{2(\text{g})} + \text{O}_{(\text{g})}$
- What is the action of selenium on magnesium metal?
- Write the name of isomerism in the following complexes:
 $[\text{Cu}(\text{NH}_3)_4]$ $[\text{PtCl}_4]$ and $[\text{Pt}(\text{NH}_3)_4]$ $[\text{CuCl}_4]$
- Write the name of the alloy used in Fischer Tropsch process in the synthesis of gasoline.
- Henry's law constant for $\text{CH}_3\text{Br}_{(\text{g})}$ is $0.159 \text{ mol dm}^{-3} \text{ bar}^{-1}$ at 25°C . What is solubility of $\text{CH}_3\text{Br}_{(\text{g})}$ in water at same temperature and partial pressure of 0.164 bar ?

SECTION – B

Attempt any EIGHT of the following questions:

[16]

- Explain pseudo-first order reaction with suitable example.
- Write the consequences of Schottky defect with reasons.
- What is the action of following on ethyl bromide:
 (i) Na in dry ether
 (ii) Mg in dry ether
- Explain formation of peptide linkage in protein with an example.
- Derive an expression to calculate molar mass of non volatile solute by osmotic pressure measurement.
- Explain monodentate and ambidentate ligands with example.
- Explain the trends in the following atomic properties of group 16 elements:
 (i) Atomic radii

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Table of Contents Hsc Chemistry Question Paper 22

1. Understanding the eBook Hsc Chemistry Question Paper 22
 - The Rise of Digital Reading Hsc Chemistry Question Paper 22
 - Advantages of eBooks Over Traditional Books
2. Identifying Hsc Chemistry Question Paper 22
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hsc Chemistry Question Paper 22
 - User-Friendly Interface
4. Exploring eBook Recommendations from Hsc Chemistry Question Paper 22
 - Personalized Recommendations
 - Hsc Chemistry Question Paper 22 User Reviews and Ratings
 - Hsc Chemistry Question Paper 22 and Bestseller Lists

5. Accessing Hsc Chemistry Question Paper 22 Free and Paid eBooks
 - Hsc Chemistry Question Paper 22 Public Domain eBooks
 - Hsc Chemistry Question Paper 22 eBook Subscription Services
 - Hsc Chemistry Question Paper 22 Budget-Friendly Options
6. Navigating Hsc Chemistry Question Paper 22 eBook Formats
 - ePub, PDF, MOBI, and More
 - Hsc Chemistry Question Paper 22 Compatibility with Devices
 - Hsc Chemistry Question Paper 22 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Hsc Chemistry Question Paper 22
 - Highlighting and Note-Taking Hsc Chemistry Question Paper 22
 - Interactive Elements Hsc Chemistry Question Paper 22
8. Staying Engaged with Hsc Chemistry Question Paper 22
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hsc Chemistry Question Paper 22
9. Balancing eBooks and Physical Books Hsc Chemistry Question Paper 22
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hsc Chemistry Question Paper 22
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Hsc Chemistry Question Paper 22
 - Setting Reading Goals Hsc Chemistry Question Paper 22
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Hsc Chemistry Question Paper 22
 - Fact-Checking eBook Content of Hsc Chemistry Question Paper 22
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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