

SECTION - I

Q. 1. Select and write the most appropriate answers from the given alternatives for each sub-question :

[7]

- (i) In Van Arkel method of refining metal, impure zirconium is converted to unstable volatile compound by heating it with –
 (a) oxygen (b) chlorine (c) bromine (d) iodine
- (ii) The temperature at which vapour pressure of a liquid becomes equal to the atmospheric pressure is _____.
 (a) melting point (b) boiling point (c) 273 K (d) 373 K
- (iii) Which of the following parameters are correct for triclinic lattice ?
 (a) $\alpha = \beta = \gamma = 90^\circ$ and $a = b = c$ (b) $\alpha \neq \beta \neq \gamma \neq 90^\circ$ and $a \neq b \neq c$
 (c) $\alpha = \gamma = 90^\circ$, $\beta \neq 90^\circ$ and $a \neq b \neq c$ (d) $\alpha \neq \beta \neq \gamma \neq 90^\circ$ and $a \neq b \neq c$
- (iv) Which mixture is used for respiration by deep sea divers ?
 (a) He + O₂ (b) Ne + O₂ (c) Ar + O₂ (d) Kr + O₂
- (v) For the reaction $2\text{N}_2\text{O}_{5(g)} \longrightarrow 4\text{NO}_{2(g)} + \text{O}_{2(g)}$ in liquid bromine, which of the following rate equation is 'incorrect'
 (a) $-\frac{1}{2} \frac{d[\text{N}_2\text{O}_5]}{dt}$ (b) $-\frac{1}{4} \frac{d[\text{NO}_2]}{dt}$ (c) $\frac{d[\text{O}_2]}{dt}$ (d) $\frac{1}{4} \frac{d[\text{NO}_2]}{dt}$
- (vi) For a certain reaction $\Delta H = -50 \text{ KJ}$ and $\Delta S = -80 \text{ JK}^{-1}$, at what temperature does the reaction turn from spontaneous to nonspontaneous.
 (a) 6.25 K (b) 62.5 K (c) 625 K (d) 6250 K
- (vii) What is the ratio of volumes of H₂ and O₂ liberated during electrolysis of acidified water ?
 (a) 1 : 2 (b) 2 : 1 (c) 1 : 8 (d) 8 : 1

Q. 2. Answer any THREE of the following :

[9]

- (i) Calculate ΔH° for the following reactions –



Given that



- (ii) Calculate molarity and molality of 6.3% solution of nitric acid having density 1.04 gcm^{-3} . (H = 1, N = 14, O = 16)
- (iii) What is the action of chlorine (Cl₂) on the following :
 (a) Cold and dilute caustic soda.
 (b) Hot and concentrated caustic soda.
 (c) Potassium bromide solution.
- (iv) Calculate the number of atoms present in 2 gram of crystal which has face-centred cubic (FCC) crystal lattice having edge length of 100 pm and density 10 gcm^{-3} .

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1. Understanding the eBook Hsc Chemistry Paper October 2013
 - The Rise of Digital Reading Hsc Chemistry Paper October 2013
 - Advantages of eBooks Over Traditional Books
2. Identifying Hsc Chemistry Paper October 2013
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hsc Chemistry Paper October 2013
 - User-Friendly Interface
4. Exploring eBook Recommendations from Hsc Chemistry Paper October 2013
 - Personalized Recommendations
 - Hsc Chemistry Paper October 2013 User Reviews and Ratings
 - Hsc Chemistry Paper October 2013 and Bestseller Lists
5. Accessing Hsc Chemistry Paper October 2013 Free and Paid eBooks

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- Hsc Chemistry Paper October 2013 eBook Subscription Services
- Hsc Chemistry Paper October 2013 Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
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 - Hsc Chemistry Paper October 2013 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Hsc Chemistry Paper October 2013
 - Highlighting and Note-Taking Hsc Chemistry Paper October 2013
 - Interactive Elements Hsc Chemistry Paper October 2013
- 8. Staying Engaged with Hsc Chemistry Paper October 2013
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hsc Chemistry Paper October 2013
- 9. Balancing eBooks and Physical Books Hsc Chemistry Paper October 2013
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hsc Chemistry Paper October 2013
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Hsc Chemistry Paper October 2013
 - Setting Reading Goals Hsc Chemistry Paper October 2013
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hsc Chemistry Paper October 2013
 - Fact-Checking eBook Content of Hsc Chemistry Paper October 2013
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

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- Interactive and Gamified eBooks

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