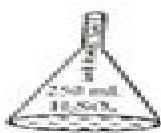


অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার



পটাসিয়াম কপ্টার, এখানে মোট পটাসিয়াম, আরও কপ্টারের শীর্ষস্থানীয় কপ্টারের পটাসিয়াম এবং মনোমের টেম্পের পটাসিয়াম
পটাসিয়াম উত্তর অধ্যায়ভিত্তিক মেট্রিক হলে। এগুলোর অনুশীলন করলে তুমি এ অধ্যায়টি থেকে সেরেফেরে পটাসিয়াম কপ্টারের
এদের উত্তর লিখতে পারবে সহজেই।

প্রশ্ন ১



পাত্র-A



পাত্র-B

সি. কে. ১০/১১/১২

- মোলের প্রবণ কী?
- "কেলস পটাসিয়াম লবণের প্রাকৃতিক, অপমাত্রা বৃষ্টির সাথে সাথে এক পর্যায়ে বৃষ্টি না পেয়ে হ্রাস পায়"— ব্যাখ্যা করো।
- A-পাত্রের বিন্যাসন এলিটের ভর কত নির্ণয় করো।
- প্রবণ দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অক্সিজেন দ্বারা পরিণত হবে— বিশ্লেষণ করো।

১ নং প্রশ্নের উত্তর

1000 mL প্রবণে কোন প্রবণের 1 mol প্রবৃত্তিত থাকলে উক্ত প্রবণকে মোলার প্রবণ বলে।

কেলস পটাসিয়াম লবণের প্রাকৃতিক অপমাত্রা বৃষ্টির সাথে সাথে এক পর্যায়ে বৃষ্টি না পেয়ে হ্রাস পায়। এর কারণ, অপমাত্রা বৃষ্টির সাথে সাথে প্রাকৃতিক বৃষ্টি পেতে থাকলেও একটি নির্দিষ্ট অপমাত্রার কেলস পটাসিয়াম লবণ নিষ্কাশিত হয়। তখন প্রাকৃতিক হ্রাস পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কথা বলা যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর ক্ষেত্রে পটাসিয়ামে প্রবৃত্তিত হওয়ার সময় বিয়োজন প্রক্রিয়াটি অপমাত্রা বৃষ্টিতে কিন্তু 32.4°C অপমাত্রার উপরে $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ যখন Na_2SO_4 এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি অপোৎপাদী হয়। এ কারণে 32.4°C অপমাত্রার উপরে অপমাত্রা বৃষ্টিজনিত কারণে প্রাকৃতিক হ্রাস ঘটে।

প্রশ্ন ২

আমরা জানি,
 $w = MV$
 $= (0.5 \times 98 \times 0.25) \text{ g}$
 $= 12.25 \text{ g}$
অতএব, প্রবৃত্তিত H_2SO_4 এর ভর
 $= 12.25 \text{ g}$

উদাহরণের A-পাত্রের অবস্থিতি,
 H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$
প্রবণের ঘনমাত্রা,
 $S = 0.5 \text{ M}$
প্রবণের আয়তন,
 $V = 250 \text{ mL}$
 $= \frac{250}{1000} = 0.25 \text{ L}$
প্রবৃত্তিত H_2SO_4 এর ভর, $w = ?$

প্রশ্ন-A: H_2SO_4 এর মোল সংখ্যা $= \left(0.5 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.125 \text{ mol}$

প্রশ্ন-B: Na_2CO_3 এর মোল সংখ্যা $= \left(0.11 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.077 \text{ mol}$

সম্প্রতিক বিক্রিয়া: $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$
সমীকরণ হতে পাই, 1 mol $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$

1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1 mol এর সাথে
 $0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol}$

অবশিষ্ট H_2SO_4 এর পরিমাণ $= (0.125 - 0.077) \text{ mol}$
 $= 0.048 \text{ mol}$

পাত্র-A এর প্রবণের সাথে পাত্র-B এর প্রবণের বিক্রিয়ার ফলে 0.048 mol H_2SO_4 অবশিষ্ট থাকায় মিশ্রিত প্রবণের প্রকৃতি হবে অম্লীয়।

প্রশ্ন ২ একমাত্র ছাত্র ল্যাবরেটরিতে X-পদ্ধতিতে 10 mg এর প্রবণের পরিমাণ এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2% $\frac{W}{V}$ HCl এর নমুনা তৈরি করে।

সি. কে. ১০/১১/১২

- নোড কাকে বলে?
- Cl_2 অপোনার— ব্যাখ্যা করো।
- উদীপকের এলিটটির মোলার ঘনমাত্রা নির্ণয় করো।
- উদীপকের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো।

২ নং প্রশ্নের উত্তর

নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেক্ট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা সমান হবে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হবে সমযোজী বন্ধন। অপরোকার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপরোকার হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও C অণুতে তড়িৎ ঋণাত্মকতার পার্থক্য $(3 - 3) = 0$, তাই Cl_2 বিশুদ্ধ সমযোজী ও অপরোকার যৌগ। উল্লেখ্য যৌগিক অণুর ক্ষেত্রে এমনিভাবে থাকে।

উদীপকে বর্ণিত এলিটটি হলো 2% $\left(\frac{W}{V}\right)$ HCl।

2% $\left(\frac{W}{V}\right)$ HCl এর অর্থ হলো—

100 mL প্রবণে HCl এর ভর 2g

∴ 1000 mL " " " " $\left(\frac{2 \times 1000}{100}\right) \text{ g}$
 $= 20 \text{ g}$

Hsc Chemistry Out Question

Jessica J Manson

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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