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**Abstract**

- ### ५. मरु-संस्कृत संज्ञा

**☞** কোলাস পানিসূত্র লবণের দ্রাব্যতা আপনাত্মা কৃষ্ণির সাথে সাথে এক পর্যায়ে কৃষ্ণি না পেয়ে হ্রাস পায়। এর কারণ, আপনাত্মা কৃষ্ণির সাথে সাথে দ্রাব্যতা কৃষ্ণি পেতে থাকলেও একটি নির্দিষ্ট আপনাত্মার কোলাস পানি দ্রুত লবণ নিষ্কৃতিত হয়। তখন দ্রাব্যতা হ্রাস পেতে থাকে। উদাহরণস্বরূপ  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  এর কথা বলা যায়।  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  এর ক্ষেত্রে পানিতে দ্রবীভূত হওয়ার সময় বিয়োজন প্রক্রিয়াটি তাপস্বরী বিক্রিয়া কিন্তু  $32.4^\circ\text{C}$  আপনাত্মার উপরে  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  যখন  $\text{Na}_2\text{SO}_4$  এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি তাপোৎপাদী হয়। এ কারণে  $32.4^\circ\text{C}$  আপনাত্মার উপরে আপনাত্মা কৃষ্ণিজনিত কারণে দ্রাব্যতা হ্রাস ঘটে।

w = SMV  
 =  $(0.5 \times 98 \times 0.25) \text{ g}$   
 = 12.25 g  
 অতএব, শুষ্কত H<sub>2</sub>SO<sub>4</sub> এর ভর  
 = 12.25 g

উদাহরণের A-পাত্রে অবস্থিত,  
 $H_2SO_4$  এর আণবিক ভর,  
 $M = 1 \times 2 + 32 + 16 \times 4$   
 $= 2 + 32 + 64$   
 $= 98 \text{ g/mol}$   
 দ্রবণের ঘনমাত্রা,  
 $S = 0.5 \text{ M}$   
 দ্রবণের আয়তন,  
 $V = 250 \text{ mL}$   
 $= \frac{250}{1000} = 0.25 \text{ L}$

ଅବଶିଷ୍ଟ  $\text{H}_2\text{SO}_4$  ଏବଂ ଭର,  $w = 7$

ক. নোড কাকে বলে?

খ.  $Cl_2$  অণুদ্বার— ব্যাখ্যা করো।

গ. উদ্দীপকের এলিভটির মৌল্যত্ব ঘনমাত্রা নির্ণয় করো।

ঘ. উদ্দীপকের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো।

२. नए व्यंजन का चिह्न

**উদাহরণ ১:** নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেকট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

১৩। সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা সম হলে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন বিশুদ্ধ ও অপোলার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য ০.৫ এর কম হলে অণুপোলার হয়। যেহেতু  $Cl_1$  এর তড়িৎ ঋণাত্মকতা ৩.০ ও  $C_2$  অণুতে তড়িৎ ঋণাত্মকতার পার্থক্য  $(3 - 3) = 0$ , তাই  $Cl_1$  বিশুদ্ধ সমযোজী ও অপোলার যৌগ। উল্লিখিত মৌলিক অণুর ক্ষেত্রে এছাড়াও থাকে।

গ) উদ্দীপকে বর্ণিত এসিডটি হলো  $2\% \left(\frac{W}{V}\right) HCl$

2%  $\left(\frac{W}{V}\right)$  HCl এর অর্থ হলো—

100 mL. 5% HCl થરા ઉપર 2g

$$\therefore 1000 \text{ mL} = \left( \frac{2 \times 1000}{100} \right) \text{ g} = 20 \text{ g}$$

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**Liyong Dong**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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### **Table of Contents Hsc Exam Chemistry Answer Paper**

1. Understanding the eBook Hsc Exam Chemistry Answer Paper
  - The Rise of Digital Reading Hsc Exam Chemistry Answer Paper
  - Advantages of eBooks Over Traditional Books
2. Identifying Hsc Exam Chemistry Answer Paper
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Hsc Exam Chemistry Answer Paper
  - User-Friendly Interface
4. Exploring eBook Recommendations from Hsc Exam Chemistry Answer Paper
  - Personalized Recommendations

- Hsc Exam Chemistry Answer Paper User Reviews and Ratings
- Hsc Exam Chemistry Answer Paper and Bestseller Lists
- 5. Accessing Hsc Exam Chemistry Answer Paper Free and Paid eBooks
  - Hsc Exam Chemistry Answer Paper Public Domain eBooks
  - Hsc Exam Chemistry Answer Paper eBook Subscription Services
  - Hsc Exam Chemistry Answer Paper Budget-Friendly Options
- 6. Navigating Hsc Exam Chemistry Answer Paper eBook Formats
  - ePub, PDF, MOBI, and More
  - Hsc Exam Chemistry Answer Paper Compatibility with Devices
  - Hsc Exam Chemistry Answer Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Hsc Exam Chemistry Answer Paper
  - Highlighting and Note-Taking Hsc Exam Chemistry Answer Paper
  - Interactive Elements Hsc Exam Chemistry Answer Paper
- 8. Staying Engaged with Hsc Exam Chemistry Answer Paper
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Hsc Exam Chemistry Answer Paper
- 9. Balancing eBooks and Physical Books Hsc Exam Chemistry Answer Paper
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Hsc Exam Chemistry Answer Paper
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Hsc Exam Chemistry Answer Paper
  - Setting Reading Goals Hsc Exam Chemistry Answer Paper
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hsc Exam Chemistry Answer Paper
  - Fact-Checking eBook Content of Hsc Exam Chemistry Answer Paper

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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