পরীকারী কপুরা, এখানে বোর্র পরীকা, ক্যাডোট কলেলগর শীর্ষান্ধানীর কলেলগমূহের পরীক্ষা এবং মহেল টেন্টের ব্যাত্তর পূর্ণাল উত্তর অধ্যায়ভিত্তিক দেবছা ময়েছে। এগুলো অনুশীলন করনে কৃষি এ অধ্যায়টি লেকে সেকোনো স্থানশীল জনামূদ্





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- ক, মোলার দ্রবণ কীয়
- "কেলাস পানিত্র লকণের দ্রাকারা, তাপমারা বৃশ্বির সাথে সাথে এক পরীয়ে বৃশ্বি না পেয়ে য়াস পায়"— ব্যাবয়া করে।
- চবশ দৃটি মিশ্রিত করলে, থিপ্রদটি কী পরিমাণ অগ্নীয় বা ফারীয় মবে— বিশ্লেষণ করো।

১ নং প্রয়ের উত্তর

1000 mL চৰণে কোন দ্ৰবের 1 mol দ্রবীভূত থাকলে উপ্ত দ্রবণকে
মোলার দ্রবন বলে।

ত্রলাস পানিত্র গবণের প্রাব্যান্তা তাপমাত্রা বৃশ্বির সাথে সাথে এক পর্যান্ত বৃশ্বির না পেরা ছাস পার। এর কারণ, তাপমাত্রা বৃশ্বির সাথে সাবে প্রব্যান্ত বৃশ্বির বিধান পার। এর কারণ, তাপমাত্রা বৃশ্বির সাথে সাবে প্রব্যান বৃশ্বির পোনে থাকার একটি নির্নিন্দ তাপমাত্রার কেলাস পানি কুর গরুণ নির্নান্ত হয়। তথান প্রাব্যান্ত হাস পোনে থাকে। উদাহরণমূল্প Na₂SO₄, 10H₂O এর কথা করা যায়। Na₂SO₄, 10H₂O এর পেরে পানিতে প্রবীল্ভ হওয়ার সময় বিয়োজন প্রক্রিয়ান্তি ভাপহারী বিক্রিয়া কিন্তু 32,4°C তাপমাত্রার উপরে Na₂SO₄, 10H₂O যখন Na₂SO₄ এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি ভাপোৎপাদী হয়। এ কারণে 32,4°C তাপমাত্রার উপরে তাপমাত্রা কৃশ্বিজনিত কারণে প্রাব্যান্ত ছাস ঘটে।

191

भ = SMV

-(0.5 × 98 × 0.25) g

= 12.25 g

অতএব, দ্বীভূত H₂SO₄ এর ভর ≒ 12.25 g

উদ্দীপকের A-পাত্রে অবস্থিত,

H₂SO₄ এর আণবিক ভর,

M = 1 × 2 + 32 + 16 × 4

= 2 + 32 + 64

= 98 g/mol

রবণের ঘনমারা,

S = 0.5 M

রবণের আরতন,

V = 250 mL

= 250

1000 = 0.25L

ফরীভূত H₂SO₄ এর ভর, w = 1

ভারত ন: H₂SO₄ এর মোল সংখ্যা — $\left(0.5 \times \frac{250}{1000}\right)$ med — 0.125 mod — 0.125 mod — 0.125 mod — 0.077 mod — 0.077 mod — 0.077 mod সংঘটিত বিভিন্ন: H₂SO₄ + Na₂CO₃ — 0.077 mod — 0.077

পাত্র-∧ এর প্রবংগর সাথে পাত্র-৪ এর প্রবংগর বিভিন্নার ফলে ১১৯ moi H₂SO₄ অবশিন্ট থাকায় মিশ্রিত প্রবংগর প্রকৃতি মধে অঞ্জীত

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ett.

ভটা ১২ একজন ছাত্র লাবরেটরিতে X-পশ্বতিতে 10 mg বর জেছ হ পরিমাণ এবং Y-পশ্বতিতে 0.5 g পরিমাণ 2% ₩ HCl এর নমূর হ কাজ করে।

- ক, নোড কাকে বলে?
- খ. Cl₂ অপোলার- ব্যাখ্যা করে।।
- উদ্দীপকের এসিভটির মোলার ঘনমারা নির্ণয় করে।
- য়, উদ্দীপকের কোন পশ্বতি পরিবেশবান্ধর? বিশ্লেষণ করে। ২ নং প্রয়োর উত্তর

নিউরিয়াসের চারনিকে যে নির্দিউ ত্রিমাত্রিক স্থানে কোন ইত্র্ক পাওয়ার সম্ভাব্যতা শূন্য তাকে নোভ বলে।

সমযোজী বন্ধনে আবন্ধ উভয় পরমাণুর তড়িং কণাছকতা ক্র হলে অর্থাৎ তড়িং কণাব্যকতার পার্থকা শূন্য হলে সমযোজী ক বিশুন্ধ ও অপোলার হয়। আর তড়িং ঝণাব্যকতার পার্থকা 0.5 এর হ মর্লে অপোলার হয়। যেহেত CI এর তড়িং ঝণাব্যকতা 3.0 ৫ ট অপুতে তড়িং ঝণাব্যকতার পার্থকা (3 – 3)= 0, তাই CI₂ ক্রিন্ সমযোজী ও অপোলার যৌগ। উরোধ্য মৌলিক অপুর ক্রেন্তে এর্ণ হা থাকে।

ত্রী উদ্দীপকে বর্ণিত এগিচটি হলো 2% $\left(\frac{W}{V}\right)$ HCI ।

2% $\left(\frac{W}{V}\right)$ HCI এর অর্থ ফলো-

.. 1000 mL $\frac{2 \times 1000}{100}$ g

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Jorge G. Ibanez, Margarita Hernandez-Esparza, Carmen Doria-Serrano, Arturo Fregoso-Infante, Mono Mohan Singh

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used as a practical tool for understanding natural phenomena and developing and improving technologies and products Concepts such as internal energy enthalpy entropy and Gibbs energy are explained using ideas and experiences familiar to students and realistic examples are given so the usefulness and pervasiveness of thermodynamics becomes apparent The worked examples illustrate key ideas and demonstrate important types of calculations and the problems at the end of chapters are designed to reinforce important concepts and show the broad range of applications Most can be solved using digitized data from open access databases and a spreadsheet Answers are provided for the numerical problems A particular theme of the book is the calculation of the equilibrium composition of systems both reactive and non reactive and this includes the principles of Gibbs energy minimization The overall approach leads to the intelligent use of thermodynamic software packages but while these are discussed and their use demonstrated they are not the focus of the book the aim being to provide the necessary foundations Another unique aspect is the inclusion of three applications chapters heat and energy aspects of processing the thermodynamics of metal production and recycling and applications of electrochemistry This book is aimed primarily at students of chemistry chemical engineering applied science materials science and metallurgy though it will be also useful for students undertaking courses in geology and environmental science A solutions manual is available for Cambridge Checkpoints HSC Chemistry 2017-19 Maureen Slade, Roger Slade, 2016-06-29 instructors Thermochemistry in Materials Processing Gabriel Plascencia, David Jaramillo, 2017-02-28 This book provides the reader with some thermochemistry notes The intention is to provide a simple easy to understand text which serves as a complimentary material to more complex books It also provide students and those beginning in the field with several application examples used in different areas of materials processing The book presents fully solved problems some quite often found in major metallurgical operations Excel Revise in a Month HSC Chemistry Will Marchment, 2007 **Research & Development** ,1993-09 Metal-Ligand Interactions in Chemistry, Physics and Biology N. Russo, Dennis R. Salahub, 2012-12-06 Proceedings of the NATO Advanced Study Institute held in Cetraro CS Italy from 1 12 September 1998 Resources in Education ,1982

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