

## SECTION - I

Q. 1. Select and write the most appropriate answer from the given alternatives for each sub-question : [7]

- (i) The hybridisation of phosphorus in phosphorus pentachloride is -  
 (a)  $dsp^3$  (b)  $sp^3d$  (c)  $d^2sp^3$  (d)  $sp^3d^2$
- (ii) The rate constant for a first order reaction is  $100\text{ S}^{-1}$ . The time required for completion of 50% of reaction is -  
 (a) 0.0693 milliseconds (b) 0.693 milliseconds  
 (c) 6.93 milliseconds (d) 69.3 milliseconds
- (iii) Silica is added to roasted copper ore during smelting process to remove -  
 (a) ferrous sulphide (b) ferrous oxide (c) cuprous sulphide (d) cuprous oxide
- (iv) 96500 coulombs correspond to the charge on how many electrons ?  
 (a)  $1.6 \times 10^{19}$  (b)  $6.022 \times 10^{20}$  (c)  $6.022 \times 10^{23}$  (d)  $6.022 \times 10^{24}$
- (v) For the reaction :  $\text{Cl}_{2(g)} \longrightarrow 2\text{Cl}_{(g)}$  -  
 (a)  $\Delta H$  is positive,  $\Delta S$  is positive (b)  $\Delta H$  is positive,  $\Delta S$  is negative  
 (c)  $\Delta H$  is negative,  $\Delta S$  is negative (d)  $\Delta H$  is negative,  $\Delta S$  is positive
- (vi) The substance 'X', when dissolved in solvent water gave molar mass corresponding to the molecular formula 'X<sub>3</sub>'. The van't Hoff factor (i) is -  
 (a) 3 (b) 0.33 (c) 1.3 (d) 1
- (vii) The relation  $a \neq b \neq c$  and  $\alpha \neq \beta \neq \gamma \neq 90^\circ$  represents which crystal system ?  
 (a) Orthorhombic (b) Tetragonal (c) Triclinic (d) Monoclinic

Q. 2. Answer any SIX of the following : [12]

- (i) Define ebullioscopic constant. Write its unit.
- (ii) State : (a) Second law of thermodynamics in terms of entropy.  
 (b) Third law of thermodynamics.
- (iii) What is the hybridisation of carbon atom in diamond and graphite ?  
 0.1 mole of Buckminster fullerene contains how many Kg of carbon ?  
 [Atomic mass of carbon = 12]
- (iv) Draw structures of (a) Chlorine trifluoride (b) Chlorine pentafluoride.

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**Sqa (Hodder Ed)**



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