

Basic Stoichiometry

Why?

In this activity we will address the question:

How do I convert between moles of different chemical species in a given reaction?

Model 1



2 mol A produces 5 mol C

3 mol B produces 4 mol D

2 mol A reacts with 3 mol B

4 mol A produces 10 mol C

6 mol B produces 8 mol D

6 mol A reacts with 9 mol B

1. What do the coefficients of a balanced equation represent?

3. If 1.5 mol B are reacted, how many mol C are formed? Show your work.

2. If 6 mol A are reacted, how many mol D are formed? Explain IN WORDS how you figured this out.

4. If 8 mol of A are reacted, how many mol B are used? Show your work.

Model 2



If 3 mol A react, how many mol C are produced?

$$3 \text{ mol A} \times \frac{5 \text{ mol C}}{2 \text{ mol A}} = 7.5 \text{ mol C}$$

$$\text{In general: } \text{mol known} \times \frac{\text{mol unknown}}{\text{mol known}} = \text{mol unknown}$$

5. If 21.0 mol A are reacted, how many mol C are produced? Using dimensional analysis, show how you calculated your answer.

6. If 21.0 mol B are reacted, how many mol A must have reacted?

Hs Chemistry Pogil Activity Basic Stoichiometry Answers

B I. Bhatt



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Calculating in Chemistry Cosmas C Mbuzi, 2022-09-20 This book is intended to help students fully grasp calculations involving reacting quantities. Students in higher courses may find it a helpful revision and enhance their clarity. The book completely discusses the key topics in basic stoichiometry including the mole concept, reacting quantities, and empirical and molecular formulas. It begins with the basic concepts and formulas required to convert various quantities to moles or amount of substance. This is particularly useful to a beginner. Chapter 2 describes how to calculate reacting quantities and therefore provides a general step by step framework or approach by which to solve these problems. The chapter also describes and applies the concept of a limiting reagent. Two methods of determining a limiting reagent are explained and illustrated. The concept of limiting reagent is extended to reacting volumes of gases with a short cut method. A short cut method for solving reacting quantities involving masses and volumes of gases is also given. Chapter 3 describes the calculations involved in the practical determination of molecular and empirical formulas. A clear meaning of percentage composition of mass is provided and used to solve problems in a step by step manner. In chapter 4 we discuss percentage yield and purity. A number of examples are given to illustrate how formulas of yield and purity are used in various circumstances. A student will find these examples helpful in relating different formulas for percentage purity. The last chapter introduces a graphical method for reacting quantities. The method may provide a new way of looking at chemical reactions. Examples are given to illustrate the method including how it can be used to determine limiting reagents. It is hoped that the book will provide all the necessary knowledge and skills to students studying an introductory chemistry course. Teachers may also find this book a good resource for their lessons in stoichiometry. *A Concrete Stoichiometry Unit for High School Chemistry* Jennifer Louise Pakkala, 2006 *A Stoichiometry Unit* David Callaghan, 2004 **Stoichiometry** Sydney Young, 1908 **Formulation and Stoichiometry** Emil J. Margolis, 2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry. As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all. The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities. For the serious student this means more often than not the frustrations of questions unanswered. The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer. Distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media, quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions, intensive application of the

universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction **Chemistry**
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