

F332 advance notice (June 2013) *"predicted Q's"*

- How does photolysis play a part in the destruction of N_2O in the stratosphere?
- Use a series of equations to show how NO acts in the catalytic destruction of ozone?
- What is meant by the term *photolysis* ?
- Convert 100 parts per billion to a percentage, for the level of NO_x in urban areas
- Why is the $\text{N}_2\text{O}_{(\text{liquid})}$ into $\text{N}_2\text{O}_{(\text{gas})}$ **not** a true dynamic equilibrium ?
- Describe how photochemical smog can be harmful to:
(a) the environment ? (b) Humans ?
- The article mentions a reaction with HO radicals, what is a radical?
Give examples of other radicals found in the stratosphere?
- Suggest why the concentration of N_2O in the atmosphere has increased since 1820 ?
- What is used as a catalyst in the upper atmosphere in photolytic fission leading to radicals being formed ?
- Explain why N_2O is considered a greenhouse gas?
- Give the symbol equation for the reaction taking place in a catalytic converter to turn nitrogen oxides (NO_x) to oxygen and nitrogen?
Is this oxidation or reduction?
- 'Nitrous oxide is released into the atmosphere from the soil aquatic systems'. Discuss?
- Why are the three bonds in nitrogen (N_2) significant ?
- Explain using examples the mechanism of a radical chain reaction?
- What are the dangers of photochemical smog?
- Explain why lightning contributes to the formation of NO_x give one of the origins of NO_x ?
- Draw a dot and cross bonding diagram for N_2O
- Loss of N_2O by wet deposition contributes to which atmospheric problem?
- Give the oxidation state of **N** in each of the following:
a) N_2O b) HNO_3 c) NO d) $(\text{NH}_4)_2\text{SO}_4$
- Use a diagram to show the bonding in Ozone (O_3)
Describe why the bonds aren't true double bonds?
- Write balanced symbol equations to represent the reactions in the use of a three-way catalyst ?
- Describe how HNO_3 is removed from that atmosphere by wet deposition ?

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