

GROUP 2 ELEMENTS AND THEIR COMPOUNDS

- show their outermost electron placed on s subshell
- readily react to lose 2 electrons
 - $\text{M} \rightarrow \text{M}^{2+} + 2\text{e}^-$

Trends down the group:

- 1st ionisation energy decreases
 - Distance between outermost electron and nucleus increases
 - More shielding
 - Counteractive effect of increasing nuclear charge
- Atomic radius increases
- Melting point generally decreases
 - Much lower than all transition metals

Physical properties:

- Good conductors of heat and electricity
 - Sea of delocalised electrons
- Form white compounds
 - Transition metals form coloured compounds
- Oxidation number of +2
- All compounds are ionic or almost ionic
 - Be compounds show more covalent character because small highly charged cation polarises anion shell of anion
- Oxides / hydroxides are basic
 - React with acids (proton acceptors)
 - Example:
 - BeO
 - $\text{BeO} + 2\text{H}^+ \rightarrow \text{Be}^{2+} + \text{H}_2\text{O}$ which is amphoteric
- Density = $2.3\text{--}3.7\text{ g/cm}^3$
 - Average density of transition metals is about 7 g/cm^3
- Compared to group 2 metals:
 - Harder + denser
 - Higher melting point
 - Metallic bond is stronger (higher charge on cation & more electrons delocalised)

Chemical properties:

- Strong reducing agents
- Oxidation is losing electrons
- Reduction is gaining electrons:
 - Reducing agent -donates electrons (reducing other substances)
 - Oxidising agent -accepts electrons (oxidising itself)
- E.g.
 - $\text{ZnO} + \text{CO} \rightarrow \text{Zn} + \text{CO}_2$
 - $\text{ZnO} \rightarrow \text{Zn}^{2+} + 2\text{O}^{2-}$
 - $\text{CO} + 2\text{O}^{2-} \rightarrow \text{CO}_2^{2-}$

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VM Jensen



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