

# Enthalpy of Formation

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- The  $\Delta H^\circ_f$  for ethanol ( $\text{C}_2\text{H}_5\text{OH}$ ) is depicted in the following thermochemical equation.



$$\Delta H^\circ_f = - 277.7 \text{ kJ/mol}$$

**Notice:** standard state of C = graphite

H =  $\text{H}_2 (\text{g})$

O =  $\text{O}_2 (\text{g})$

can use fractional coefficients

# Enthalpy Change Of Formation Value For $\text{KHCO}_3$

**Thomas Griffiths**



### Enthalpy Change Of Formation Value For $\text{K}_2\text{CO}_3$ :

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