

# Measurements Using Electrochemical Cells and Electroplating

The basic counting unit in chemistry, the mole, has a special name, *Avogadro's number*, in honor of the Italian scientist Amadeo Avogadro (1776-1856). The commonly accepted definition of Avogadro's number is the number of atoms in exactly 12 g of the isotope  $^{12}\text{C}$ , and the quantity itself is  $6.02214199(47) \times 10^{23}$ .<sup>1</sup>

A bit of information about Avogadro seems appropriate. His full name was Lorenzo Romano Amedeo Carlo Avogadro (almost a mole of letters in his name). He was a practicing lawyer until 1806 when he began his new career teaching physics and math at the University of Turin, where he was later promoted to the chair of physical chemistry. In 1811, Avogadro published a paper in the *Journal de Physique*, entitled "Essay on a Manner of Determining the Relative Masses of the Elementary Molecules of Bodies, and the Proportions in Which They Enter into These Compounds," which pretty much says it all. This paper includes the statement that has come to be regarded as Avogadro's Hypothesis:

The first hypothesis to present itself in this connection, and apparently even the only admissible one, is the supposition that the number of integral molecules in any gases is always the same for equal volumes, or always proportional to the volumes. Indeed, if we were to suppose that the number of molecules contained in a given volume were different for different gases, it would scarcely be possible to conceive that the law regulating the distance of molecules could give in all cases relations as simple as those which the facts just detailed compel us to acknowledge between the volumes and the number of molecules.

In this experiment, you will confirm Avogadro's number by conducting an electrochemical process called *electrolysis*. In electrolysis, an external power supply is used to drive an otherwise *nonspontaneous* reaction. You will use a copper strip and a zinc strip as the electrodes, placed in a beaker of sulfuric acid. You will make the cell electrolytic by using the copper strip as the anode and the zinc strip as the cathode. By determining the average current used in the reaction, along with the knowledge that all of the copper ions formed are the  $2^+$  cations, you will calculate the number of atoms in one mole of copper and compare this value with Avogadro's number.

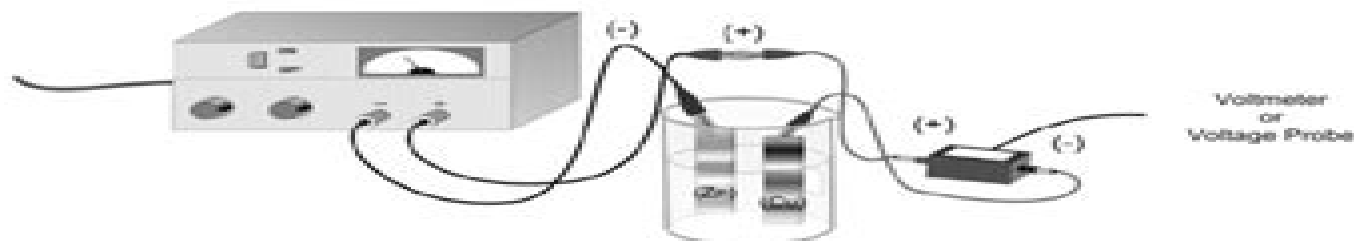


Figure 1

<sup>1</sup> CRC Handbook of Chemistry and Physics, 82<sup>nd</sup> edition, p. 1-7.

# Measurements Using Electrochemical Cells And Electroplating

**LM Reder**



## **Measurements Using Electrochemical Cells And Electroplating:**

The book delves into Measurements Using Electrochemical Cells And Electroplating. Measurements Using Electrochemical Cells And Electroplating is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Measurements Using Electrochemical Cells And Electroplating, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:
    - Chapter 1: Introduction to Measurements Using Electrochemical Cells And Electroplating
    - Chapter 2: Essential Elements of Measurements Using Electrochemical Cells And Electroplating
    - Chapter 3: Measurements Using Electrochemical Cells And Electroplating in Everyday Life
    - Chapter 4: Measurements Using Electrochemical Cells And Electroplating in Specific Contexts
    - Chapter 5: Conclusion
  2. In chapter 1, the author will provide an overview of Measurements Using Electrochemical Cells And Electroplating. This chapter will explore what Measurements Using Electrochemical Cells And Electroplating is, why Measurements Using Electrochemical Cells And Electroplating is vital, and how to effectively learn about Measurements Using Electrochemical Cells And Electroplating.
  3. In chapter 2, this book will delve into the foundational concepts of Measurements Using Electrochemical Cells And Electroplating. This chapter will elucidate the essential principles that must be understood to grasp Measurements Using Electrochemical Cells And Electroplating in its entirety.
  4. In chapter 3, the author will examine the practical applications of Measurements Using Electrochemical Cells And Electroplating in daily life. This chapter will showcase real-world examples of how Measurements Using Electrochemical Cells And Electroplating can be effectively utilized in everyday scenarios.
  5. In chapter 4, the author will scrutinize the relevance of Measurements Using Electrochemical Cells And Electroplating in specific contexts. The fourth chapter will explore how Measurements Using Electrochemical Cells And Electroplating is applied in specialized fields, such as education, business, and technology.
  6. In chapter 5, this book will draw a conclusion about Measurements Using Electrochemical Cells And Electroplating. This chapter will summarize the key points that have been discussed throughout the book.
- The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Measurements Using Electrochemical Cells And Electroplating.

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