

1 Consider the isoelectronic species, Na^+ , Mg^{2+} , F^- , and O^{2-} . The correct order of increasing length of their radii is

1. $\text{F}^- < \text{O}^{2-} < \text{Mg}^{2+} < \text{Na}^+$
2. $\text{Mg}^{2+} < \text{Na}^+ < \text{F}^- < \text{O}^{2-}$
3. $\text{O}^{2-} < \text{F}^- < \text{Na}^+ < \text{Mg}^{2+}$
4. $\text{O}^{2-} < \text{F}^- < \text{Mg}^{2+} < \text{Na}^+$

2 Among the following elements, the one that is not an actinoid is:

1. Curium ($Z=96$)
2. Californium ($Z=98$)
3. Uranium ($Z=92$)
4. Terbium ($Z=65$)

3 The order of screening effect of electrons of s, p, d and f orbitals of a given shell of an atom on electrons in its outer shell is :

1. $s > p > d > f$
2. $f > d > p > s$
3. $p < d < f < s$
4. $f > p > s > d$

4 The first ionisation enthalpies of Na, Mg, Al, and Si are in the order of-

1. $\text{Na} < \text{Al} < \text{Mg} < \text{Si}$
2. $\text{Na} > \text{Mg} > \text{Al} > \text{Si}$
3. $\text{Na} < \text{Mg} < \text{Al} < \text{Si}$
4. $\text{Na} > \text{Mg} > \text{Al} < \text{Si}$

5 Among halogens, the correct order of amount of energy released in electron gain (electron gain enthalpy) is-

1. $\text{F} > \text{Cl} > \text{Br} > \text{I}$
2. $\text{F} < \text{Cl} < \text{Br} < \text{I}$
3. $\text{F} < \text{Cl} > \text{Br} > \text{I}$
4. $\text{F} < \text{Cl} < \text{Br} < \text{I}$

6 The period number in the long form of the periodic table is equal to :

1. Magnetic quantum number of any element of the period.
2. Atomic number of any element of the period.
3. Maximum principal quantum number of any element of the period.
4. Maximum azimuthal quantum number of any element of the period.

7 The correct order of size of the given species is:

1. $\text{I} > \text{I}^- > \text{I}^+$
2. $\text{I}^+ > \text{I}^- > \text{I}$
3. $\text{I} > \text{I}^+ > \text{I}^-$
4. $\text{I}^- > \text{I} > \text{I}^+$

8 Electronic configuration of four elements A, B, C and D are given below :

- A. $1s^2 2s^2 2p^3$
 B. $1s^2 2s^2 2p^4$
 C. $1s^2 2s^2 2p^6 3s^1$
 D. $1s^2 2s^2 2p^5$

The correct order of increasing tendency to gain electron is :

1. $\text{A} < \text{C} < \text{B} < \text{D}$
2. $\text{A} < \text{B} < \text{C} < \text{D}$
3. $\text{D} < \text{B} < \text{C} < \text{A}$
4. $\text{D} < \text{A} < \text{B} < \text{C}$

9 Match the correct ionization enthalpies and electron gain enthalpies of the following elements.

	Elements		ΔH_1	ΔH_2	$\Delta_{\text{eg}} H$
(i)	Most reactive non-metal	A.	419	3051	-48
(ii)	Most reactive metal	B.	1681	3374	-328
(iii)	Least reactive element	C.	738	1451	-40
(iv)	Metal forming binary halide	D.	2372	5251	+48

Codes

	A	B	C	D
1.	ii	i	iv	iii
2.	i	ii	iii	iv
3.	i	iv	iii	ii
4.	iv	i	iii	ii

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 - The Rise of Digital Reading Mcqs On Inorganic
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcqs On Inorganic
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcqs On Inorganic
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcqs On Inorganic
 - Personalized Recommendations
 - Mcqs On Inorganic User Reviews and Ratings

- Mcqs On Inorganic and Bestseller Lists
- 5. Accessing Mcqs On Inorganic Free and Paid eBooks
 - Mcqs On Inorganic Public Domain eBooks
 - Mcqs On Inorganic eBook Subscription Services
 - Mcqs On Inorganic Budget-Friendly Options
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 - Mcqs On Inorganic Enhanced eBook Features
- 7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking Mcqs On Inorganic
 - Interactive Elements Mcqs On Inorganic
- 8. Staying Engaged with Mcqs On Inorganic
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcqs On Inorganic
- 9. Balancing eBooks and Physical Books Mcqs On Inorganic
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcqs On Inorganic
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mcqs On Inorganic
 - Setting Reading Goals Mcqs On Inorganic
 - Carving Out Dedicated Reading Time
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 - Fact-Checking eBook Content of Mcqs On Inorganic
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
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