

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Period: \_\_\_\_\_

### Midterm Exam Study Guide

#### d. Isotopes

- Isotopes differ in their \_\_\_\_\_ and number of \_\_\_\_\_.

#### e. Average atomic mass Formula:

#### f. Ions

- \_\_\_\_\_ have lost electrons and have a \_\_\_\_\_ charge
- \_\_\_\_\_ have gained electrons and have a \_\_\_\_\_ charge

#### g. Group Numbers be able to locate them

- Alkali Metals
- Alkaline earth Metals
- Transition Metals
- Inner Transition Metals
- Halogens
- Noble Gases

#### h. Periodic Trends Give your notes for reference (the charts for each of these)

- Atomic Radius/Size
- Ionic Radius/Size
- Ionization Energy
- Electronegativity

#### i. Valence Electrons be able to identify # of valence electrons and charge for each group

- Lewis Dot structures

### F. Nuclear Decay and Half Life

- $^0_{-1}e$  or  $^0_{-1}\beta$  is called an \_\_\_\_\_ particle which has the weakest radiation.
- $^0_{+1}e$  or  $^0_{+1}\beta$  is called an \_\_\_\_\_ particle which has moderate radiation and is an electron.
- $^4_2\alpha$  is called an \_\_\_\_\_ particle which has the strongest radiation and highest energy.
- Half-life
  - Definition: \_\_\_\_\_
  - Know how to read a graph and calculate half-life

### d. Electron Configurations

- Be able to draw them/identify them from another student
- Hund's Rule
- Pauli Exclusion Principle
- Aufbau Principle

### e. Ionic Compounds

- Ionic compounds are made of \_\_\_\_\_ + \_\_\_\_\_
- Their overall charge is \_\_\_\_\_
- Know how to use Lewis Dot Structures to draw bonds between cation and anion

# Mcps Chemistry Semester Review Guide Answers

**Zhe Xu**



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**Resources in Education** ,1985      **Subject Guide to Books in Print** ,1971      **Dissertation Abstracts**

**International** ,2004      *Chemistry - the Physical Setting* ,2004      **AP Chemistry Exam Success** Lewis Morris,2018-07-14

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*AP Chemistry Test Secrets*, 2010

*Student Study Guide/Solutions Manual to accompany General, Organic & Biological Chemistry* Erin Smith Berk, Janice Smith, 2012-02-10 Each chapter of the Student Study Guide begins with a chapter review tied to the chapter goals in the text Next Sample problems are supplied and stepped out through the solution for each type of problem covered in the chapter A Self Test serves up fill in the blank exercises to assess learning with answers supplied at the end of the chapter Finally chapters end with the solutions for all of the in chapter problems as well as for the odd numbered end of chapter problems

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