

Like Dissolves Like Practice Answers

Using your knowledge of the "like dissolves like" principle of solubility, determine which component will dissolve best in each of the following solvents. Explain your reason for choosing each like substance.

1) Solvent: methanol (CH₃OH)

Components: carbon tetrachloride and oxygen dichloride

Both oxygen dichloride and methanol are polar, while carbon tetrachloride is not.

2) Solvent: carbon dioxide

Components: methanol (CH₃OH) and nitrogen dioxide

While both components are polar, methanol is considerably more polar than nitrogen dioxide (both components always have that charge separation, while nitrogen dioxide, on certain times polar, the rest).

3) Solvent: water

Components: carbon tetrachloride and ammonia

Both water and ammonia are polar. Carbon tetrachloride is not.

4) Solvent: hexagonal alcohol

Components: water and methanol

Both water and hexagonal alcohol are polar. Methanol is not.

Like Dissolves Like Practice Answer Key

Daniel F McAuley



Like Dissolves Like Practice Answer Key:

Unveiling the Magic of Words: A Review of "**Like Dissolves Like Practice Answer Key**"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is really awe-inspiring. Enter the realm of "**Like Dissolves Like Practice Answer Key**," a mesmerizing literary masterpiece penned by way of a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve to the book is central themes, examine its distinctive writing style, and assess its profound affect the souls of its readers.

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