

(ix) The standard potential of the cell in the following reaction is _____



($E_{\text{Cd}^{2+}}^{\circ} = -0.403 \text{ V}$, $E_{\text{Cu}^{2+}}^{\circ} = 0.334 \text{ V}$)

(a) -0.737 V

(b) 0.737 V

(c) -0.069 V

(d) 0.069 V

(x) The value of $[\text{H}_3\text{O}^+]$ in mol lit^{-1} of 0.001 M acetic acid solution ($K_a = 1.8 \times 10^{-5}$) is _____

(a) 1.34×10^{-1}

(b) 1.34×10^{-2}

(c) 1.34×10^{-3}

(d) 1.34×10^{-4}

Q.2. Answer the following questions:

[8]

(i) Write the product formed when alkyl halide reacts with silver nitrite.

(ii) Write the name of product formed, when acetone is treated with 2, 4-dinitrophenyl hydrazine.

(iii) Write the name of biodegradable polyamide copolymer.

(iv) Identify the molecularity of following elementary reaction:



(v) What is the action of selenium on magnesium metal?

(vi) Write the name of isomerism in the following complexes:



(vii) Write the name of the alloy used in Fischer Tropsch process in the synthesis of gasoline.

(viii) Henry's law constant for $\text{CH}_3\text{Br}_{(\text{g})}$ is $0.159 \text{ mol dm}^{-3} \text{ bar}^{-1}$ at 25°C . What is solubility of $\text{CH}_3\text{Br}_{(\text{g})}$ in water at same temperature and partial pressure of 0.164 bar ?

SECTION – B

Attempt any EIGHT of the following questions:

[16]

Q.3. Explain pseudo-first order reaction with suitable example.

Q.4. Write the consequences of Schottky defect with reasons.

Q.5. What is the action of following on ethyl bromide:

(i) Na in dry ether

(ii) Mg in dry ether

Q.6. Explain formation of peptide linkage in protein with an example.

Q.7. Derive an expression to calculate molar mass of non volatile solute by osmotic pressure measurement.

Q.8. Explain monodentate and ambidentate ligands with example.

Q.9. Explain the trends in the following atomic properties of group 16 elements:

(i) Atomic radii

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Table of Contents Hsc Board Question Papers Science 22

1. Understanding the eBook Hsc Board Question Papers Science 22
 - The Rise of Digital Reading Hsc Board Question Papers Science 22
 - Advantages of eBooks Over Traditional Books
2. Identifying Hsc Board Question Papers Science 22
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hsc Board Question Papers Science 22
 - User-Friendly Interface
4. Exploring eBook Recommendations from Hsc Board Question Papers Science 22

- Personalized Recommendations
 - Hsc Board Question Papers Science 22 User Reviews and Ratings
 - Hsc Board Question Papers Science 22 and Bestseller Lists
5. Accessing Hsc Board Question Papers Science 22 Free and Paid eBooks
- Hsc Board Question Papers Science 22 Public Domain eBooks
 - Hsc Board Question Papers Science 22 eBook Subscription Services
 - Hsc Board Question Papers Science 22 Budget-Friendly Options
6. Navigating Hsc Board Question Papers Science 22 eBook Formats
- ePub, PDF, MOBI, and More
 - Hsc Board Question Papers Science 22 Compatibility with Devices
 - Hsc Board Question Papers Science 22 Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Hsc Board Question Papers Science 22
 - Highlighting and Note-Taking Hsc Board Question Papers Science 22
 - Interactive Elements Hsc Board Question Papers Science 22
8. Staying Engaged with Hsc Board Question Papers Science 22
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hsc Board Question Papers Science 22
9. Balancing eBooks and Physical Books Hsc Board Question Papers Science 22
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hsc Board Question Papers Science 22
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Hsc Board Question Papers Science 22
- Setting Reading Goals Hsc Board Question Papers Science 22
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Hsc Board Question Papers Science 22

- Fact-Checking eBook Content of Hsc Board Question Papers Science 22
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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