

Experiment 1

1. Write the general equation for an equilibrium/phase transition and give the conditions of each. Example: melting

Experiment 2: Liquid and Vapor Equilibrium in a Closed System

In a closed system, equilibrium is established between the liquid and the vapor of a substance. The conditions of equilibrium are:

- a. pressure is equal to the external pressure
b. the temperature is equal to the boiling point

a.	b.
b.	a.

Experiment 3: Vapor Pressure of a Pure Substance

1. The vapor pressure of a pure substance is the pressure exerted by the vapor of the substance in equilibrium with the liquid phase.

- a. What is the relationship between the vapor pressure and the temperature?

- b. What is the relationship between the vapor pressure and the boiling point?

- c. What is the relationship between the vapor pressure and the density?

- d. What is the relationship between the vapor pressure and the heat of vaporization?

2. The vapor pressure of a pure substance is a function of temperature. The Clausius-Clapeyron equation can be used to calculate the vapor pressure of a pure substance at any temperature.

- a. What is the Clausius-Clapeyron equation?

- b. What is the Clausius-Clapeyron equation?

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Annelies Wilder-Smith



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