

ChemThink: Behavior of Gases

Name _____ Per _____

1. What does the speed of atoms depend on? Describe the relationship between speed and each of these factors.

temperature: $\uparrow$ temp = $\uparrow$ speed

type of atom: smaller/lighter = faster

2. In a sample of one type of atom, do all the atoms move at the same speed? Why?

No, speed is related to each individual atoms energy

3. What is the equation for Pressure and what are the units for pressure?

$$\text{Pressure} = \frac{\text{Force}}{\text{Area}} \quad \text{psi}$$

4. When atoms are moving in a closed container what is creating the force on the sides of the container?

impacts (collisions) of atoms w/ sides of container

5. Explain how a change in temperature causes pressure to change and include an equation.

more collisions results in higher pressure

$\downarrow$
b/c molecules moving faster

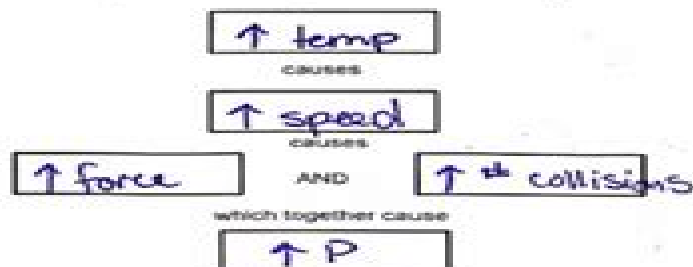
$$\frac{P}{T} = \text{constant}$$

6. Draw the graph you observed of pressure vs. temperature. Is this a direct relationship or an inverse relationship?



direct

7. Fill the cause/effect relationship in the diagram of boxes you observed.



- Explain how the number of atoms present causes pressure to change and include an equation?

more atoms = more collisions, which increases pressure

$$\frac{P}{n} = \text{constant}$$

How To Do Behavior Of Gases Chem Think

American Chemical Society

The logo of the American Chemical Society, featuring a red semi-circle with a white border.

How To Do Behavior Of Gases Chem Think:

Introduction to Chemical Structure Donald Cotter,2020-08-01 **School Science and Mathematics** ,1918

Chemistry in the Community (ChemCom) American Chemical Society,2011-06-17 Touted as the most successful NSF funded project published Chemistry in the Community ChemCom by the American Chemical Society ACS offers a meaningful and memorable chemistry program for all levels of high school students ChemCom covers traditional chemistry topics within the context of societal issues and real world scenarios Centered on decision making activities where students are responsible for generating data in an investigating analyzing that data and then applying their chemistry knowledge to solve the presented problem The text is intensively laboratory based with all 39 of the investigations integrated within the text not separate from the reading With the ChemCom program students learn more organic and biochemistry more environmental and industrial chemistry and more on the particulate nature of matter than other textbooks all within the relevance of solving problems that arise in everyday life Meticulously updated to meet the needs of today s teachers and students the new sixth edition of ChemCom adheres to the new science framework as well as the forthcoming next generation of science standards Incorporating advances in learning and cognitive sciences ChemCom s wide ranging coverage builds upon the concepts and principles found in the National Science Education Standards Correlations are available showing how closely aligned ChemCom is to these and other state standards Chemistry Carla Mooney,2016-05-16 Have you ever wondered what makes up everything in the world around you Or what exactly is the difference between solids liquids and gases Have you wanted to know what causes two substances to react or change Chemistry Investigate the Matter that Makes Up Your World introduces readers 12 through 15 to the fascinating world of protons neutrons and electrons Learn how these molecules combine to form ordinary objects such as the chair you re sitting on the water in your glass even you Through hands on investigative projects readers delve into the world of chemical reactions and changing matter learning how these principles are used in many areas of science from biochemistry to nuclear science Combining hands on science inquiry with chemistry mathematics and biology projects include building models of molecules and bonds identifying acids and bases investigating the effect of temperature on reaction rate and observing how a chemical reaction from vinegar water and bleach can accelerate the rusting of steel Chemistry offers entertaining illustrations and fascinating sidebars to illuminate the topic and engage readers further plus integrates a digital learning component by providing links to primary sources videos and other relevant websites *From Elements to Atoms* Robert Siegfried,2002 Seeking to enlarge an understanding of the nature of chemical science references to recently published secondary works are intended for the benefit of readers who wish to enlarge their perspectives on the development of early chemical thinking *More Chemistry Basics* William C. Robertson,2010 Overwhelmed by orbitals Terrified of thermodynamics Agitated by acids and bases Have no fear This follow up to the award winning Chemistry Basics will clear up your chemistry woes In More Chemistry Basics the ninth book in the

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An Introduction to Chemistry Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life. It provides a major overview of topics typically found in first year chemistry courses in the USA. The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box. Numerous examples are presented in every chapter to aid students and provide helpful self learning tools. The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry.

The Freudian Paradigm Mohammed Mujeeb-ur-Rahman, 1977 *American Gas-light Journal and Chemical Repertory*, 1909 *Physical and Chemical Mechanisms in Molecular Radiation Biology* William A. Glass, Matesh N. Varma, 2012-12-06

The fundamental understanding of the production of biological effects by ionizing radiation may well be one of the most important scientific objectives of mankind. Such understanding could lead to the effective and safe utilization of the nuclear energy option. In addition, this knowledge will be of immense value in such diverse fields as radiation therapy and diagnosis and in the space program. To achieve the above stated objective, the U.S. Department of Energy (DOE) and its predecessors embarked upon a fundamental interdisciplinary research program some 35 years ago. A critical component of this program is the Radiological and Chemical Physics Program (RCPP). When the RCPP was established, there was very little basic knowledge in the fields of physics, chemistry and biology that could be directly applied to understanding the effects of radiation on biological systems. Progress of the RCPP program in its first 15 years was documented in the proceedings of a conference held at Airlie, Virginia in 1972. At this conference, it was clear that considerable progress had been made in research on the physical and chemical processes in well characterized systems that could be used to understand biological effects. During this period of time, most physical knowledge was obtained for the gas phase because the technology and instrumentation had not progressed to the point that measurements could be made in liquids more characteristic of biological materials.

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