

Name

Key

1. Hydrogen, H_2 , is used as rocket fuel. The hydrogen is burned in oxygen, O_2 , according to the following equation. What is the enthalpy change if 200.0 g of hydrogen, H_2 , is burned?



$$\frac{200.0 \text{ g } H_2}{2.016 \text{ g } H_2} \times \frac{1 \text{ mol } H_2}{1} = 99.2 \text{ mol } H_2$$

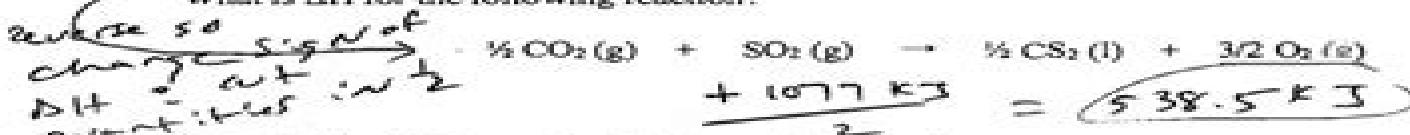
$$99.2 \text{ mol } H_2 \times \frac{-484 \text{ kJ}}{2 \text{ mol } H_2} = -24,006 \text{ kJ}$$

-24,006 kJ

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



What is ΔH for the following reaction?



3. Following is the equation for the combustion of acetone:



How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

$$\frac{-853.6 \text{ kcal}}{2 \text{ mol } C_3H_6O} \times 4.5 \text{ mol } C_3H_6O = -1920.6 \text{ kcal}$$

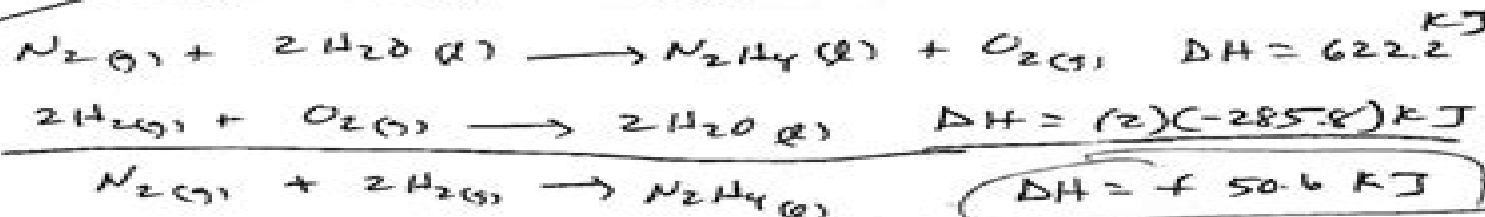
Exothermic

4. The specific heat of iron is $0.110 \text{ cal} / \text{g} \cdot ^\circ\text{C}$. How much heat must be added to 5.2 g of iron at 25°C to raise its temperature to 225°C ?

$$\text{heat} = q = m c \Delta T = (5.2 \text{ g}) \left(\frac{0.110 \text{ cal}}{\text{g} \cdot ^\circ\text{C}} \right) (225 - 25^\circ\text{C}) = 114.4 \text{ cal}$$

Endothermic

5. Hydrazine, N_2H_4 , is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate ΔH for reaction (1) given the data below.



General Chemistry Practice Problems With Answers

Daniela Niemeyer



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Table of Contents General Chemistry Practice Problems With Answers

1. Understanding the eBook General Chemistry Practice Problems With Answers
 - The Rise of Digital Reading General Chemistry Practice Problems With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Practice Problems With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a General Chemistry Practice Problems With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Practice Problems With Answers
 - Personalized Recommendations
 - General Chemistry Practice Problems With Answers User Reviews and Ratings
 - General Chemistry Practice Problems With Answers and Bestseller Lists

5. Accessing General Chemistry Practice Problems With Answers Free and Paid eBooks
 - General Chemistry Practice Problems With Answers Public Domain eBooks
 - General Chemistry Practice Problems With Answers eBook Subscription Services
 - General Chemistry Practice Problems With Answers Budget-Friendly Options
6. Navigating General Chemistry Practice Problems With Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - General Chemistry Practice Problems With Answers Compatibility with Devices
 - General Chemistry Practice Problems With Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of General Chemistry Practice Problems With Answers
 - Highlighting and Note-Taking General Chemistry Practice Problems With Answers
 - Interactive Elements General Chemistry Practice Problems With Answers
8. Staying Engaged with General Chemistry Practice Problems With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry Practice Problems With Answers
9. Balancing eBooks and Physical Books General Chemistry Practice Problems With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Chemistry Practice Problems With Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine General Chemistry Practice Problems With Answers
 - Setting Reading Goals General Chemistry Practice Problems With Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of General Chemistry Practice Problems With Answers
 - Fact-Checking eBook Content of General Chemistry Practice Problems With Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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