

Chemistry 105 Equation Sheet

Constants

$$\begin{aligned}a_0 &= 0.5291772059 \text{ \AA} \\c &= 2.99792458 \times 10^8 \text{ m/s} \\e &= 1.6021765 \times 10^{-19} \text{ C} \\\epsilon_0 &= 8.8541877 \times 10^{-12} \text{ F/m} \\\hbar &= 6.626 \times 10^{-34} \text{ J}\cdot\text{s} \\k_B &= 1.38065 \times 10^{-23} \text{ J/K} \\k_w &= 8.987551 \times 10^8 \frac{\text{N}\cdot\text{m}^2}{\text{C}^2} \\m_e &= 9.109383 \times 10^{-31} \text{ kg} \\m_n &= 1.6749273 \times 10^{-27} \text{ kg} \\m_p &= 1.6726217 \times 10^{-27} \text{ kg} \\N_A &= 6.022142 \times 10^{23} \text{ mol}^{-1} \\R &= 8.31447 \frac{\text{J}}{\text{mol}\cdot\text{K}} \\&= 0.0820575 \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}} \\&= 1.987207 \frac{\text{cal}}{\text{mol}\cdot\text{K}} \\R_H &= 1.0974 \times 10^7 \text{ m}^{-1}\end{aligned}$$

Conversions (* = exact)

Length

$$\begin{aligned}1 \text{ in} &= 2.54 \text{ cm}^* \\1 \text{ ft} &= 12 \text{ inches}^* \\1 \text{ yd} &= 3 \text{ ft}^* \\1 \text{ m} &= 39.37 \text{ inches}^* \\1 \text{ mile} &= 5280 \text{ feet}^* \\1 \text{ mile} &= 1.609 \text{ km}\end{aligned}$$

Mass

$$\begin{aligned}1 \text{ lb} &= 453.59237 \text{ g} \\&= 16 \text{ oz. avdp.}^* \\1 \text{ US lb} &= 12 \text{ troy ounces}^* \\1 \text{ ton} &= 2000 \text{ lb}^* \\1 \text{ amu} &= 1.660538 \times 10^{-24} \text{ g}\end{aligned}$$

Volume

$$\begin{aligned}1 \text{ US gal} &= 3.7854 \text{ L} = 4 \text{ US qt}^* \\1 \text{ US qt} &= 32 \text{ US fl. oz.}^* \\1 \text{ US fluid oz.} &= 29.5735 \text{ mL} \\1 \text{ cm}^3 &= 0.001 \text{ L}^* \\1 \text{ in}^3 &= 16.4 \text{ cm}^3\end{aligned}$$

Pressure

$$\begin{aligned}1 \text{ atm} &= 760 \text{ torr}^* \\&= 101325 \text{ Pa}^* \\&= 406.783 \text{ in H}_2\text{O} \\&= 14.6959 \text{ lb/in}^2 \text{ (psi)} \\1 \text{ torr} &= 1 \text{ mm Hg (at } 0^\circ\text{C)}^* \\1 \text{ Pa} &= 1 \text{ kg}\cdot\text{m}^{-1}\cdot\text{s}^{-2} \\1 \text{ bar} &= 100,000 \text{ Pa}^*\end{aligned}$$

Force

$$1 \text{ N} = 1 \text{ kg}\cdot\text{m}\cdot\text{s}^{-2}$$

Power

$$1 \text{ Watt} = 1 \text{ J/s}$$

Energy

$$\begin{aligned}1 \text{ J} &= 1 \text{ kg}\cdot\text{m}^2/\text{s}^2 = 0.2390\text{-cal} \\1 \text{ J} &= 1 \text{ V}\cdot\text{C} = 0.009369 \text{ L}\cdot\text{atm} \\1 \text{ eV} &= 1.602 \times 10^{-19} \text{ J} \\1 \text{ L}\cdot\text{atm} &= 101.3 \text{ J}\end{aligned}$$

Formulas

$$\begin{aligned}\rho &= m/V \\T_F &= \frac{9}{5}T_C + 32^\circ\text{F} \\T_K &= \frac{1}{1}T_C + 273.15 \text{ K} \\c &= \lambda \nu \\E &= h\nu \\\lambda &= \frac{h}{mv} \\E_{\text{kinetic}} &= \frac{1}{2}mv^2 \\E &= mc^2 \\\Phi &= h\nu - KE_{\text{electron}} \\E_n &= -Z^2 \frac{R_H h c}{n^2}\end{aligned}$$

$$\Delta x \Delta(mv) \geq \frac{h}{4\pi}$$

$$\frac{1}{\lambda} = R_H \left(\frac{1}{n_1^2} - \frac{1}{n_2^2} \right)$$

$$V(r) = \frac{q_1 q_2}{4\pi\epsilon_0 r}$$

$$F(r) = k_e \frac{Q_1 Q_2}{r^2}$$

$$\mu = Qr$$

$$\Delta E = q + w$$

$$\Delta E_{\text{univ}} = 0$$

$$\Delta S_{\text{univ}} \geq 0$$

$$S = k_B \ln W$$

$$q = mC\Delta T$$

$$w = -P_{\text{ext}} \Delta V$$

$$\Delta H = \Delta E + P\Delta V$$

$$\Delta S = \frac{q_{\text{rev}}}{T}$$

$$\Delta G = \Delta H - T\Delta S$$

$$\Delta G = \Delta G^\circ + RT \ln Q$$

$$\Delta H'_{\text{rxn}} = \sum nH'_{f,\text{prod}} - \sum nH'_{f,\text{react}}$$

$$\Delta H'_{\text{rxn}} \approx \sum nD_{\text{broken}} - \sum nD_{\text{formed}}$$

$$PV = nRT$$

$$P_{\text{tot}} = P_1 + P_2 + P_3 + \dots$$

$$\frac{\text{Rate}_A}{\text{Rate}_B} = \frac{\text{Time}_B}{\text{Time}_A} = \sqrt{\frac{M_B}{M_A}}$$

$$KE_{\text{ave}} = \frac{3}{2}RT$$

$$u_{\text{rms}} = \sqrt{v^2} = \sqrt{\frac{3RT}{M}}$$

$$pX = -\log X$$

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-]$$

* = exact ν = frequency v = velocity

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General Chemistry Formula And Equations Sheet

M Walker



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General Chemistry Workbook Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material **MCAT General Chemistry Review** Princeton Review (Firm), 2010 The MCAT is a test of more than just the facts about basic physical and biological sciences it s an in depth rigorous examination of your knowledge of scientific concepts and principles as well as your critical thinking and writing skills With the Princeton Review s subject specific MCAT series you can focus your review on the MCAT topics that are most challenging to you Each book in the series contains the most in depth coverage of subjects tested on the MCAT Each chapter in MCAT General Chemistry includes Full color illustrations charts and diagrams Examples of general chemistry questions and their solutions worked out step by step Chapter Review Quizzes and answers A real MCAT style practice passage with questions and answers Bulleted chapter summaries for quick review MCAT General Chemistry Review also includes A complete glossary of general chemistry terms A general chemistry formula sheet **Laboratory Experiments in General Chemistry** Charles Robert Sanger, 1896

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briefing complete right leo fuel tanks quantity check throle 1000 rpm rwy in use winds ldg distance wing top surface check
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november 26 2006 data plate engine on engine make and model serial number cfr45 11 21 182 yes no n a 10 data plates
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nacelle group as follows engine section for visual evidence of excessive oil fuel or hydraulic leaks and sources of such leaks
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