

- Given the partial pressure of N_2 is 0.75 atm and the K_{a1} at 25 °C and in water is 6.7×10^{-4} what is the concentration of N_2 in water?
- At 25 °C $K_c = 7.9 \times 10^2$ for the reaction: $2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$. What is the value of K_c for the reaction: $2SO_3(g) \rightleftharpoons 2SO_2(g) + O_2(g)$?

- Consider the following reactions and their K_c values at 25 °C:



What is the equilibrium constant for the reaction: $PCl_5(g) + 2NO_2(g) \rightleftharpoons 2NOCl(g) + PCl_3(g)$ in terms of K_1 and K_2 ?

- Given the following pair of reactions and their equilibrium constants:



Calculate a value for K_c for the reaction: $NOCl(g) + \frac{1}{2} O_2(g) \rightleftharpoons NO(g) + \frac{1}{2} Cl_2(g)$

- Given the following pair of reactions and their equilibrium constants:



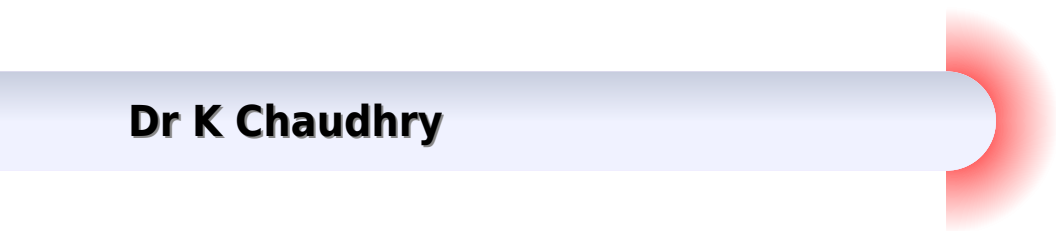
What is equilibrium constant for the reaction?



- The equilibrium constant for the reaction: $H_2(g) + I_2(g) \rightleftharpoons 2HI(g)$ is 54.9 at 699 K. What is the equilibrium constant for the reaction $4HI(g) \rightleftharpoons 2H_2(g) + 2I_2(g)$ under the same conditions?
- Given the reaction, $2NO_2(g) + O_2(g) \rightleftharpoons 2NO(g)$, where $\Delta H^\circ = -110.4 \text{ kJ/mol}$. Which one of the following will cause an increase in the equilibrium concentration of NO in a closed reaction chamber?
 - adding some more $O_2(g)$ through an injection needle
 - increasing the temperature of the system
 - removing the $NO_2(g)$ from the system
 - increasing the pressure of the system while the temperature is kept constant
 - adding a catalyst
- The reaction, $2SO_2(g) \rightleftharpoons 2SO_3(g) + O_2(g)$ is endothermic. Predict what will happen if the temperature is increased.
- Consider the following system, which is at equilibrium: $CO_2(g) + 3H_2(g) \rightleftharpoons CH_4(g) + H_2O(g)$. What would be the result of removing $CH_4(g)$ and $H_2O(g)$?

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