

Practice Test Ch 3

Stoichiometry

Name Kelly # 10

- Remember, this is practice - Do NOT cheat yourself of finding out what you are capable of doing. Be sure you follow the testing conditions outlined below.
- DO NOT USE A CALCULATOR. AP Chem does not allow the use of a calculator for the MC part of the exam, so it is time to start practicing without one. You may use ONLY a periodic table.
- While you should practice working as fast as possible, it is more important at this point in the course, that you practice without a calculator, even if it slows you down. Look for the "easy math" - common factors and rough estimation - do not do "long division" to try to get exact values. Remember it is a MC test, use the answers.
- Mark which questions you would like to "go over" when we get to school in September.

1. Balance the following equation:



The balanced equation shows that 1.00 mole of  $\text{NH}_3$  requires \_\_\_ mole(s) of  $\text{O}_2$ .

- a. 0.57
- b. 1.25
- c. 1.33
- d. 1.75
- e. 3.5

$$1 \text{ mol NH}_3 \left( \frac{7 \text{ mol O}_2}{4 \text{ mol NH}_3} \right) =$$

2. Write a balanced equation for the combustion of acetaldehyde,  $\text{CH}_3\text{CHO}$ .



When properly balanced, the equation indicates that \_\_\_ mole(s) of  $\text{O}_2$  are required for each mole of  $\text{CH}_3\text{CHO}$ .

- a. 1
- b. 2
- c. 2.5
- d. 3
- e. 6

$$1 \text{ mol CH}_3\text{CHO} \left( \frac{5 \text{ mol O}_2}{2 \text{ mol CH}_3\text{CHO}} \right) =$$

3. What is the total mass of products formed when 16 grams of  $\text{CH}_4$  is burned with excess oxygen?

- a. 32 g
- b. 36 g
- c. 44 g
- d. 62 g
- e. 80 g

$$\text{CH}_4 + 2 \text{O}_2 \rightarrow \text{CO}_2 + 2 \text{H}_2\text{O}$$

$$16 \text{ g CH}_4 \left( \frac{1 \text{ mol CH}_4}{16 \text{ g CH}_4} \right) \left( \frac{2 \text{ mol O}_2}{1 \text{ mol CH}_4} \right) \left( \frac{32 \text{ g O}_2}{1 \text{ mol O}_2} \right) =$$

4. Write a balanced equation for the combustion of propane,  $\text{C}_3\text{H}_8$ .



When balanced, the equation indicates that \_\_\_ moles of  $\text{O}_2$  are required for each mole of  $\text{C}_3\text{H}_8$ .

- a. 1.5
- b. 3
- c. 2.5
- d. 5
- e. 8

5. Balance the following equation with the SMALLEST WHOLE NUMBER COEFFICIENTS possible. Select the number that is the sum of the coefficients in the balanced equation:



- a. 3
- b. 5
- c. 6
- d. 7
- e. 8

6. Calculate the mass of hydrogen formed when 27 g of aluminum reacts with excess hydrochloric acid according to the balanced equation below:



- a. 1.5 g
- b. 2.0 g
- c. 3.0 g
- d. 6.0 g
- e. 12 g

$$27 \text{ g Al} \left( \frac{1 \text{ mol Al}}{27 \text{ g Al}} \right) \left( \frac{3 \text{ mol H}_2}{2 \text{ mol Al}} \right) \left( \frac{2 \text{ g H}_2}{1 \text{ mol H}_2} \right) =$$

7. How many grams of nitric acid,  $\text{HNO}_3$ , can be prepared from the reaction of 138 g of  $\text{NO}_2$  with 54.0 g  $\text{H}_2\text{O}$  according to the equation below?



- a. 92
- b. 108
- c. 126
- d. 189
- e. 279

$$138 \text{ g NO}_2 \left( \frac{1 \text{ mol NO}_2}{46 \text{ g NO}_2} \right) \left( \frac{2 \text{ mol HNO}_3}{3 \text{ mol NO}_2} \right) \left( \frac{63 \text{ g HNO}_3}{1 \text{ mol HNO}_3} \right) = 126 \text{ g}$$

$$54 \text{ g H}_2\text{O} \left( \frac{1 \text{ mol H}_2\text{O}}{18 \text{ g H}_2\text{O}} \right) \left( \frac{2 \text{ mol HNO}_3}{1 \text{ mol H}_2\text{O}} \right) \left( \frac{63 \text{ g HNO}_3}{1 \text{ mol HNO}_3} \right) = 4(63) \text{ g}$$

8. Which of the following statements is true?

- I. The molar mass of  $\text{CaCO}_3$  is  $100.1 \text{ g mol}^{-1}$ .
  - II. 50 g of  $\text{CaCO}_3$  contains  $9 \times 10^{23}$  oxygen atoms.
  - III. A 200 g sample of  $\text{CaCO}_3$  contains 2 moles of  $\text{CaCO}_3$ .
- a. I only
  - b. II only
  - c. III only
  - d. I and III only
  - e. I, II, and III

$$\frac{40}{12} + \frac{48}{100.1} \text{ mol}$$

# General Chemistry Stoichiometry Test Answers

**Peterson's**



## General Chemistry Stoichiometry Test Answers:

*General Chemistry Workbook* Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

**MCAT General Chemistry Review 2023-2024** Kaplan Test Prep, 2022-08-02 Kaplan's MCAT General Chemistry Review 2023 2024 offers an expert study plan detailed subject review and hundreds of online and in book practice questions all authored by the experts behind the MCAT prep course that has helped more people get into medical school than all other major courses combined Prepping for the MCAT is a true challenge Kaplan can be your partner along the way offering guidance on where to focus your efforts and how to organize your review This book has been updated to match the AAMC's guidelines precisely no more worrying about whether your MCAT review is comprehensive The Most Practice More than 350 questions in the book and access to even more online more practice than any other MCAT general chemistry book on the market The Best Practice Comprehensive general chemistry subject review is written by top rated award winning Kaplan instructors Full color 3 D illustrations from Scientific American charts graphs and diagrams help turn even the most complex science into easy to visualize concepts All material is vetted by editors with advanced science degrees and by a medical doctor Online resources including a full length practice test help you practice in the same computer based format you'll see on Test Day Expert Guidance High yield badges throughout the book identify the topics most frequently tested by the AAMC We know the test The Kaplan MCAT team has spent years studying every MCAT related document available Kaplan's expert psychometricians ensure our practice questions and study materials are true to the test

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*MCAT Exam Prep* Bill T Reese, 2024-05-13 What does it take to embark on the journey towards a career in medicine Dive into the world of aspiring medical professionals with our comprehensive guide designed to equip you with the knowledge skills and strategies needed to conquer the Medical College Admission Test MCAT and pave the way for success in medical school We unravel the mysteries of the MCAT through a series of engaging chapters that delve into the core subjects tested on the exam From biology and biochemistry to physics and organic chemistry each chapter offers a deep exploration of essential concepts accompanied by practice questions and detailed explanations to reinforce your understanding But mastering the MCAT is not just about memorizing facts and formulas it's about honing your critical thinking analytical reasoning and problem solving abilities That's why our book goes beyond mere content review to provide valuable insights into effective study strategies test taking techniques and mental preparation tips to help you perform at your best on exam day Whether you're a pre med student embarking on your MCAT journey or a seasoned test taker seeking to improve your scores this book is your trusted companion every step of the way Join us as we empower you to embrace the challenges seize the opportunities and embark on a transformative journey towards realizing your dreams of a career in medicine Prepare to embark on an enriching and empowering voyage one that will not only shape your academic future but also ignite your passion for healing compassion and making a meaningful difference in the world Are you ready to rise to the challenge and unlock your potential for greatness The journey begins here

*AP Chemistry Premium, 2024: 6 Practice Tests + Comprehensive Review + Online*

*Practice* Neil D. Jespersen, Pamela Kerrigan, 2023-07-04 A guide to taking the Advanced Placement exam in chemistry featuring a review of major chemistry concepts practice and diagnostic tests test taking strategies an overview of the test and practice problems

**SAT Subject Test Chemistry** Joseph A. Mascetta, Mark Kernion, 2020-12-01 Note College Board has discontinued the SAT Subject Tests in the US The tests will be available outside the US in June 2021 and then be discontinued Barron's SAT Subject Test Chemistry with 7 Practice Tests features in depth review of all topics on the exam and full length practice tests in the book and online This edition includes One full length diagnostic test to help you assess your strengths and weaknesses Comprehensive review of all topics on the exam including introductory chemistry atomic structure and the periodic table bonding chemical formulas gases and laws stoichiometry liquids solids and phase changes chemical reactions and thermochemistry chemical reactions chemical equilibrium acids bases and salts oxidation reduction carbon and organic chemistry and the laboratory Four full length practice tests that reflect the actual SAT Subject Test Chemistry exam in length question types and degree of difficulty Two full length online practice tests with answer explanations and automated scoring Appendices which include the periodic table important equation constant and data tables and a glossary of chemistry terms

**General Chemistry, Study Guide** James E. Brady, 1990-03-01 The Fifth Edition retains the pedagogical strengths that made the previous editions so popular and has been updated reorganized and streamlined Changes include more accessible introductory chapters with greater stress on the logic of the periodic table earlier introduction of redox reactions greater emphasis on the concept of energy a new section on Lewis structures earlier introduction of the ideal gas law and a new development of thermodynamics Each chapter ends with review questions and problems

**Exploring General Chemistry in the Laboratory** Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science

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by chemistry education are critically discussed highlighting the pitfalls that can occur in teaching chemistry and how to circumvent them The main topics discussed include best practices project based education blended learning and the role of technology including e learning and science visualization Hands on recommendations on how to optimally implement innovative strategies of teaching chemistry at university and high school levels make this book an essential resource for anybody interested in either teaching or learning chemistry more effectively from experience chemistry professors to secondary school teachers from educators with no formal training in didactics to frustrated chemistry students

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