

Unit 1: General Chemistry II study Guide

Chapter 12: Liquids, Solids, and Intermolecular Forces

Definitions:

- intermolecular forces: attractive forces that exist between all molecules and atoms
- thermal energy: energy associated with the random motion caused by increasing temperature
- crystalline solids: solids that have atoms and molecules arranged in a well-ordered 3D array
- amorphous: has atoms or molecules with no long-range orders
- dispersion force: the only intermolecular force present in all molecules and atoms, aka London dispersion forces, result from fluctuations in the electron distribution within molecules or atoms
- miscibility: the ability to mix without separating into two states
- hydrogen bonding: intermolecular force exhibited by polar molecules containing hydrogen atoms bonded directly to small electronegative atoms, primarily F, O, or N
- hydrogen bond: strong attraction between H atom on one molecule and the F, O, or N on its neighbor
- chemical bonds: occur between individual atoms within a molecule
- ion dipole forces: occur when an ionic compound is mixed with a polar compound
- surface tension: energy required to increase the surface area (of a liquid) by a unit amount
- viscosity: resistance of a liquid to flow
- capillary action: ability of a liquid to flow against gravity up a narrow tube
- vaporization: transition from a liquid to a gas
- condensation: transition from gas to a liquid, opposite of vaporization
- volatile: liquids that vaporize easily
- nonvolatile: liquids that don't vaporize easily
- heat of vaporization: amount of heat required to vaporize one mole of a liquid to gas
- vapor pressure: pressure of a gas in dynamic equilibrium with its liquid
- dynamic equilibrium: when the rate of the forward and reverse reaction are equal
- boiling point: temperature at which the liquid's vapor pressure equals the external pressure
- normal boiling point: temperature at which its vapor pressure equals 1 atm
- critical temperature: temperature at which the meniscus between the liquid and the gas disappears and the gas and liquid states commingle to form a supercritical fluid
- critical pressure: pressure at which this transition occurs
- sublimation: transition from a solid to a gas
- deposition: opposite of sublimation
- heat of fusion: amount of heat required to melt 1 mol of a solid
- triple point: point in a phase diagram that represents the unique set of conditions at which the three states are equally stable and in equilibrium
- critical point: in a phase diagram represents the temperature and pressure above which a supercritical fluid exists

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KJ Lindholm-Leary



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