

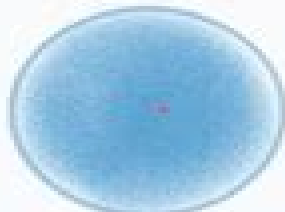
ELECTRON PROBABILITIES OF HYDROGEN ATOMS

HYDROGEN ATOM A



H
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HYDROGEN ATOM B

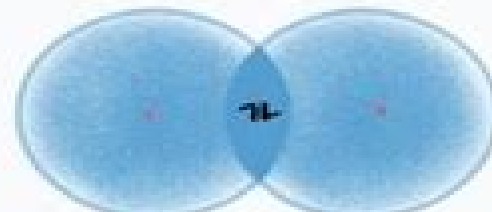


H
1s

FORMATION OF
COVALENT BOND
→

ELECTRON PROBABILITY OF HYDROGEN (H_2) MOLECULE

HYDROGEN (H_2) MOLECULE



H:H

SHARING ELECTRON DENSITY:
INCREASING DENSITY BETWEEN ATOMS

COVALENT BONDS

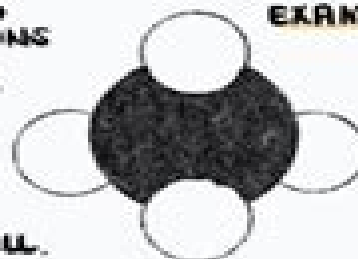
BONDING INTERACTIONS BETWEEN NONMETALS

WHEN DISCUSSING IONIC BONDING, WE STUDIED HOW THE DONATION AND ACCEPTANCE OF ELECTRONS RESULTED IN AN ELECTROSTATIC INTERACTION RESULTING IN AN IONIC BOND BETWEEN METALS AND NONMETALS. HOWEVER, COVALENT BONDS ARE FORMED THROUGH AN ENTIRELY DIFFERENT METHOD - SHARING ELECTRONS. THIS SHARING OF ELECTRONS IS FORMED BETWEEN ATOMS WITH HIGH IONIZATION ENERGIES (NONMETALS), SO NEITHER CONTENDER WANTS TO LOSE ELECTRONS FROM THEIR VALENCE SHELL. INSTEAD, THE TWO ATOMS SHARE THEIR VALENCE ELECTRONS BY MERGING THEIR UNFILLED ORBITALS, PAIRING THE VALENCE ELECTRONS THROUGHOUT THE FORMATION OF THE BOND. THE TWO ATOMS ARE ATTRACTED TO ONE ANOTHER SINCE THE BONDING ELECTRONS SIT BETWEEN THE TWO NUCLEI - THIS ATTRACTION FORCE PULLS THE ATOMS CLOSER, LOWERING THE POTENTIAL ENERGY THROUGH THE FORMATION OF THE BOND.

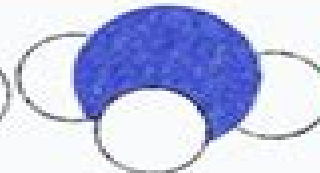
THROUGH SHARING ELECTRONS, NEITHER ELECTRON HAS TO ABSORB ENERGY TO RELEASE AN ELECTRON, AND BOTH BONDING ATOMS ARE FURTHER STABILIZED BY FILLING THEIR ORBITALS (INCREASING THE SYMMETRY OF THEIR VALENCE SHELL).

FOR EXAMPLE, H_2 IS A MOLECULE FORMED FROM A COVALENT BOND OF TWO HYDROGEN ATOMS BY BONDING THE SINGLE UNPAIRED ELECTRONS FROM EACH ATOM (WITH OPPOSING SPIN STATES - PAULI EXCLUSION PRINCIPLE).

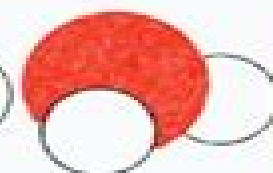
EXAMPLE OF COVALENT MOLECULES



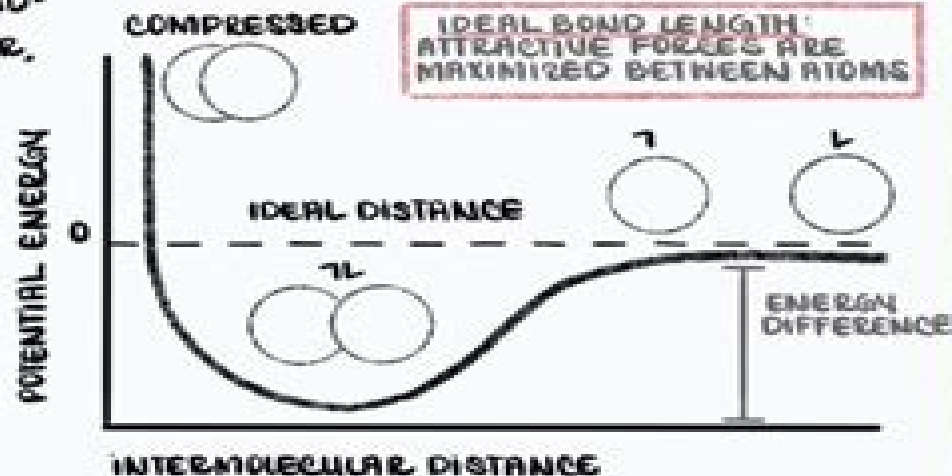
METHANE (CH_4)



AMMONIA (NH_3)



WATER (H_2O)



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