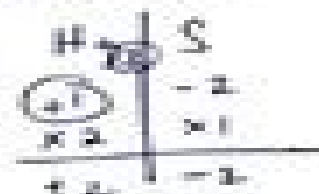
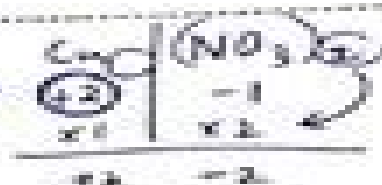


Worksheet #3: Naming Inorganic Compounds

1. Give the name for each of the following ionic compounds:

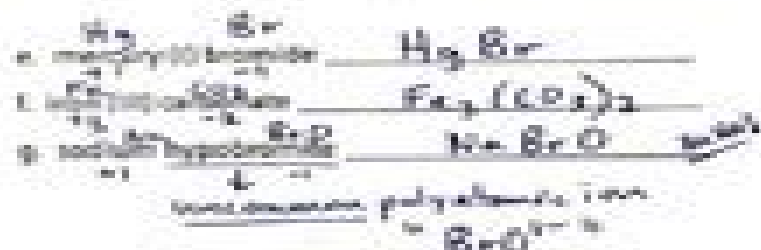
- AlF_3 Aluminum Fluoride
- Fe(OH)_2 Iron(II) Hydroxide
- $\text{Cu(NO}_3)_2$ Copper(II) Nitrate
- $\text{Ba(ClO}_3)_2$ Barium Perchlorate
- Li_3PO_4 Lithium Phosphate
- HgS Mercury(II) Sulfide
- $\text{Ca(CH}_3\text{COO)}_2$ Calcium Acetate
- $\text{Cr}_2(\text{CO}_3)_3$ Chromium(III) Carbonate
- $\text{K}_2\text{Cr}_2\text{O}_7$ Potassium Dichromate
- $\text{(NH}_4)_2\text{SO}_4$ Ammonium Sulfate



all of the underlined compounds are polyatomic ions → Memorize Them

2. Write the chemical formula for each of the following compounds:

- copper(II) oxide Cu_2O
- potassium peroxide K_2O_2
- aluminum hydroxide Al(OH)_3
- iron(III) bromide FeBr_3



3. Give the name or chemical formula, as appropriate, for each of the following acids:

- nitric acid HNO_3
- hydrobromic acid HBr
- phosphoric acid H_3PO_4

- hypochlorous acid HClO
- iodic acid HIO_3
- sulfurous acid H_2SO_3

Band off of Iodate is IO_3^{1-}

4. Give the name or chemical formula, as appropriate, for each of the following molecular substances:

- iodine hexafluoride IF_6
- iodine pentafluoride IF_5
- Xenon Trioxide XeO_3

- dinitrogen tetroxide N_2O_4
- hydrogen cyanide HCN
- tetraphosphorus hexasulfide P_4S_6

There are all binary ionic compounds (+/-)

0 - None
1 - One
2 - Two
3 - Three
4 - Four
5 - Five
6 - Six
7 - Seven
8 - Eight
9 - Nine

Gen Chem Unit 3 Packet Answers

WJ Hussar



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