

balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds:

- a. CuSO_4 copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \rightarrow 20 \rightarrow 10 \rightarrow 5 \rightarrow 2.5 \rightarrow 1.25$$

1 2 3 4 5 61.5 yr

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$100 \rightarrow 50 \rightarrow 25 \rightarrow 12.5 \text{ mg}$$

1 2 3 12.5 mg

3 half lives

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 \times .692 & = & 43.596 \rightarrow 44 \\ 65 \times .308 & = & 20.02 \rightarrow 20. \\ \hline & + & \\ & 63.62 & \rightarrow 64 \end{array}$$

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 - The Rise of Digital Reading Gen Chem 15final Exam Review Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Gen Chem 15final Exam Review Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gen Chem 15final Exam Review Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gen Chem 15final Exam Review Guide
 - Personalized Recommendations
 - Gen Chem 15final Exam Review Guide User Reviews and Ratings
 - Gen Chem 15final Exam Review Guide and Bestseller Lists

5. Accessing Gen Chem 15final Exam Review Guide Free and Paid eBooks
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 - Gen Chem 15final Exam Review Guide Enhanced eBook Features
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 - Highlighting and Note-Taking Gen Chem 15final Exam Review Guide
 - Interactive Elements Gen Chem 15final Exam Review Guide
8. Staying Engaged with Gen Chem 15final Exam Review Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gen Chem 15final Exam Review Guide
9. Balancing eBooks and Physical Books Gen Chem 15final Exam Review Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gen Chem 15final Exam Review Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gen Chem 15final Exam Review Guide
 - Setting Reading Goals Gen Chem 15final Exam Review Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gen Chem 15final Exam Review Guide
 - Fact-Checking eBook Content of Gen Chem 15final Exam Review Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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