

Equation Sheet for A-Level Chemistry

These are the equations you need to

memorise for your Chemistry A-Level



Amount of Substance

Moles

$$n = \frac{m}{M_r}$$

Concentration

$$c = \frac{n}{V}$$

Molar gas volume

$$V = n \times 24$$

Number of particles

$$= n \times N_A$$

Ideal gas Equation

$$pV = nRT$$

Atom economy

$$= \frac{\text{M}_r \text{ of desired product}}{\text{Sum of M}_r \text{ of all reactants}} \times 100$$

% Yield

$$= \frac{\text{Actual yield}}{\text{Theoretical yield}} \times 100$$

Energetics

Enthalpy change

$$Q = mc\Delta T$$

Average bond enthalpy

$$\Delta H = \sum (\text{bond enthalpies in reactants}) - \sum (\text{bond enthalpies in products})$$

Electrode Potentials

EMF

$$E^{\ominus} = E^{\ominus} \text{right hand side} - E^{\ominus} \text{left hand side}$$

Thermodynamics

Gibbs Free Energy Equation

$$\Delta G = \Delta H - T\Delta S$$

Transition Metals

Coloured complex ions - difference in energy

$$\Delta E = h\nu = \frac{hc}{\lambda}$$

Chemical Equilibria

Equilibrium constant wrt concentration

$$K_c = \frac{[D]^d [E]^e}{[A]^a [B]^b}$$

Equilibrium constant wrt pressure

$$K_p = \frac{(p_D)^d (p_E)^e}{(p_A)^a (p_B)^b}$$

Rate Equations

$$\text{Rate} = k[A]^m[B]^n$$

Arrhenius is provided but you should be able to use:

$$k = Ae^{-\frac{E_a}{RT}}$$

$$\ln k = -\frac{E_a}{RT} + \ln A$$

Acids and Bases

$$pH = -\log_{10} [H^+]$$

$$K_w = [H^+][OH^-]$$

$$pK_a = -\log_{10} K_a$$

$$K_a = \frac{[H^+][A^-]}{[HA]}$$

Mcmuc A Level Chemistry

Katrin Zwirglmaier



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Table of Contents Mcmuc A Level Chemistry

1. Understanding the eBook Mcmuc A Level Chemistry
 - The Rise of Digital Reading Mcmuc A Level Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcmuc A Level Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcmuc A Level Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcmuc A Level Chemistry
 - Personalized Recommendations
 - Mcmuc A Level Chemistry User Reviews and Ratings
 - Mcmuc A Level Chemistry and Bestseller Lists
5. Accessing Mcmuc A Level Chemistry Free and Paid eBooks
 - Mcmuc A Level Chemistry Public Domain eBooks
 - Mcmuc A Level Chemistry eBook Subscription Services
 - Mcmuc A Level Chemistry Budget-Friendly Options
6. Navigating Mcmuc A Level Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Mcmuc A Level Chemistry Compatibility with Devices
 - Mcmuc A Level Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mcmuc A Level Chemistry
 - Highlighting and Note-Taking Mcmuc A Level Chemistry
 - Interactive Elements Mcmuc A Level Chemistry
8. Staying Engaged with Mcmuc A Level Chemistry

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Mcmuc A Level Chemistry
- 9. Balancing eBooks and Physical Books Mcmuc A Level Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcmuc A Level Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mcmuc A Level Chemistry
 - Setting Reading Goals Mcmuc A Level Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mcmuc A Level Chemistry
 - Fact-Checking eBook Content of Mcmuc A Level Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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