

- The chemistry of water

- The two hydrogens are covalently bonded the oxygen (polar covalent bonds)
- The properties of water arise from attractions between oppositely charged atoms of different water molecules- the partially positive hydrogen of one molecule attracts the partially negative oxygen of a nearby molecule
- Has hydrogen bond- intermolecular forces
- When water is in its liquid state, the hydrogen bonds are very fragile

- Four properties of water that contribute to Earth's suitability for life

1. Cohesion of water molecules

- Water molecules stay close to one another as a result of hydrogen bonding
- Liquid water is constantly changing, with their molecules being rearranged- these linkages (hydrogen bonding) make water more structured than most other liquids, with the hydrogen bonds collectively holding the substance together and this is known as **cohesion**
- this result is high surface tension
- things such as spiders take advantage of this, allowing it to walk across a pond
- cohesion also contributes towards the transport of water and dissolved nutrients against gravity in plants, with the water being transported from the roots to the leaves via a network of water-conducting cells.

- As water evaporates from a leaf, the hydrogen bonds causes water molecules leaving the veins to pull on molecules farther down

- Adhesion: the clinging of one substance to another
- Adhesion of water by hydrogen bonds to the molecules of the cell walls (of the plant) help counteract the pull of gravity

2. Moderation of temperature by water

- Water molecules moderate air temperature by absorbing heat from the air that is warmer and releasing stored heat to the air when it becomes cooler, thus we refer to water as a heat bank.

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