

2015

Subject : Chemistry (Honours)

Paper : V

Time: 2 Hours

Full Marks: 50

*The figures in the margin indicate full marks.
Candidates are required to give their answers in their own words
as far as practicable.*

1. Answer any four questions:

2×4=8

- Depict the ambidentate and bidentate behaviours of thiocyanate anion.
- SiCl_4 has lower boiling point than CCl_4 . Explain.
- Cite and draw one example each of $3c-2e$ and $2c-2e$ bonds.
- Briefly comment on the compound K_3C_{60} .
- Write down the structural formulae of sodium pyrophosphate and sodium metaphosphate.
- What is meant by chalcogenide? Give example.
- 'NaOH is base but ClOH is acid' - Give reason.

2. Answer any three questions:

6×3=18

- What are inner metallic complexes? Give example. 2
 - Chromium(III) chloride forms three different hydrates of the same mole ratio $\text{Cr:Cl:H}_2\text{O}=1:3:6$. The violet form does not lose water over conc. H_2SO_4 and gives 3 equivalent of AgCl on treatment with AgNO_3 . The other two forms are green and lose 1 and 2 moles of H_2O over conc. H_2SO_4 and give 2 and 1 equivalent of AgCl , respectively with AgNO_3 . Write down the structures of the three isomeric coordination compounds and explain the isomerism involved. 4
- Define and exemplify chelate complex. 3
 - 'The stability of $[\text{Ni}(\text{en})_3]^{2+}$ (en =ethylenediamine) is much greater than that of $[\text{Ni}(\text{NH}_3)_6]^{2+}$, although both contain six Ni-N bonds.' - Explain. 3
- Apply radius ratio principle to explain the difference between crystal structures of NaCl and CsCl . 2
 - ' MgO is a stable ionic solid though large amount of energy is required to form Mg to Mg^{2+} and to dissociate O_2 into atomic oxygen and thereafter to form O^{2-} ion.' - Account for the fact. 2

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