

- According to the Bronsted—Lowry definition, which chemical species can function *both* as an acid and as a base?
 - Cl^-
 - SO_4^{2-}
 - NH_4^+
 - HCO_3^-
 - H_3O^+
- In the reaction $\text{CN}^- + \text{H}_2\text{O} \rightleftharpoons \text{HCN} + \text{OH}^-$ which is an acid—base conjugate pair?
 - H_2O and HCN
 - H_2O and OH^-
 - CN^- and H_2O
 - HCN and OH^-
- Given that HX is a stronger Bronsted acid than HY in aqueous solution, which is true of a 1 M solution of NaX ?
 - It is less basic than a 1 M solution of NaY .
 - It is more basic than a 1 M solution of NaY .
 - It yields a neutral solution.
 - It is more concentrated than a 1 M solution of NaY .
- The pH of a 0.03 M HCl solution is
 - 1.5
 - 2.5
 - 3.5
 - 12.5
- The pH of a 1.0×10^{-2} M $\text{Ba}(\text{OH})_2$ solution at 25°C is
 - 2.7
 - 3.0
 - 11.0
 - 11.3
- Which is the strongest acid?
 - HClO
 - HClO_2
 - HClO_3
 - HClO_4
- Which salt reacts with water (hydrolyzes) to produce a basic solution?
 - $\text{NaC}_2\text{H}_3\text{O}_2$
 - NaNO_3
 - NH_4Cl
 - BaSO_4

- The *weakest* of the bases listed is

Acid	Conjugate Base	K_a (Ionization Constant of Acid)
HCl	Cl^-	100% ionized
HSO_4^-	SO_4^{2-}	1.2×10^{-2}
H_2S	HS^-	5.7×10^{-8}
HS^-	S^{2-}	1.2×10^{-13}

 - Cl^-
 - CN^-
 - HS^-
 - S^{2-}
 - SO_4^{2-}

- The oxide of which element will react with water to form the strongest acid?

	Main Groups							
Period	I	II	III	IV	V	VI	VII	O
First								
Second	X	Y	Z	P		Q	S	U
Third	W					R	T	M

 - W
 - M
 - P
 - R
 - Z

- Which statement is a logical inference from the fact that a 0.10 M solution of potassium acetate $\text{KC}_2\text{H}_3\text{O}_2$ is less alkaline than a 0.10 M solution of potassium cyanide, KCN ?
 - Hydrocyanic acid is a weaker acid than acetic acid.
 - Hydrocyanic acid is less soluble in water than acetic acid.
 - Cyanides are less soluble than acetates.
 - Acetic acid is a weaker acid than hydrocyanic acid.
- In the titration of 50.0 mL of 0.100 M benzoic acid (a monoprotic acid) with 50.0 mL of 0.100 M NaOH , the properties of the solution at the equivalence point will correspond exactly to the properties of
 - a 0.100 M sodium solution.
 - a 0.0500 M sodium hydroxide solution.
 - a 0.0500 M benzoic acid solution.
 - a 0.0500 M sodium benzoate solution.

General Chemistry Acs Exam

John K. Gilbert



General Chemistry Acs Exam:

The book delves into General Chemistry Acs Exam. General Chemistry Acs Exam is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into General Chemistry Acs Exam, encompassing both the fundamentals and more intricate discussions.

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- Chapter 2: Essential Elements of General Chemistry Acs Exam
- Chapter 3: General Chemistry Acs Exam in Everyday Life
- Chapter 4: General Chemistry Acs Exam in Specific Contexts
- Chapter 5: Conclusion

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